

1. Stoichiometry is based on the law of conservation of

- | | | | |
|-------------------------|--------|-------------------------|-----------|
| <input type="radio"/> A | charge | <input type="radio"/> B | volume |
| <input type="radio"/> C | mass | <input type="radio"/> D | reactants |

2. Mole ratios are obtained from the

- | | | | |
|-------------------------|----------------------------|-------------------------|----------------|
| <input type="radio"/> A | balanced chemical equation | <input type="radio"/> B | periodic table |
| <input type="radio"/> C | molar mass | | |

3. How many particles are in 13.5 grams of Beryllium?

- | | | | |
|-------------------------|---------------|-------------------------|------------------------------|
| <input type="radio"/> A | 9 particles | <input type="radio"/> B | 4×10^{23} particles |
| <input type="radio"/> C | 1.5 particles | <input type="radio"/> D | 9×10^{23} particles |

4. $2 \text{NaClO}_3 (\text{s}) \rightarrow 2 \text{NaCl} (\text{s}) + 3 \text{O}_2 (\text{g})$
12.00 moles of NaClO_3 will produce how many grams of O_2 ?

- | | | | |
|-------------------------|-----------------------|-------------------------|-----------------------|
| <input type="radio"/> A | 288 g O_2 | <input type="radio"/> B | 576 g of O_2 |
| <input type="radio"/> C | 256 g of O_2 | <input type="radio"/> D | 598 g O_2 |

5. When 12 moles of O_2 reacts with 1.1 mole of C_{10}H_8 what is the limiting reactant? $\text{C}_{10}\text{H}_8 + 12 \text{O}_2 \rightarrow 10 \text{CO}_2 + 4 \text{H}_2\text{O}$

- | | | | |
|-------------------------|----------------|-------------------------|---------------------------|
| <input type="radio"/> A | Oxygen | <input type="radio"/> B | C_{10}H_8 |
| <input type="radio"/> C | Carbon Dioxide | <input type="radio"/> D | Water |

6. $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$ How much aluminum would be needed to completely react with 45 grams of O_2 ?

- | | | | |
|-------------------------|-------------|-------------------------|------------|
| <input type="radio"/> A | 1.875 moles | <input type="radio"/> B | 1.05 moles |
| <input type="radio"/> C | 1.875 grams | <input type="radio"/> D | 3.75 moles |

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Kaplan Test Prep



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Formulation and Stoichiometry Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction **The Role of**

Non-Stoichiometry in the Functional Properties of Oxide Materials Maria Veronica Ganduglia-Pirovano, Javier Carrasco, Claudio Cazorla, 2019-12-09 Admission Assessment Exam Review E-Book HESI, 2016-02-04 Passing your admission assessment exam is the first step on the journey to becoming a successful health professional make sure you are prepared with Admission Assessment Exam Review 4th Edition From the testing experts at HESI this user friendly guide walks you through the topics and question types found on admission exams including math reading comprehension vocabulary grammar biology chemistry anatomy and physiology and physics The guide includes hundreds of sample questions step by step explanations illustrations and comprehensive practice exams to help you review the subject areas and hone your test taking skills Plus the pre test and post test help identify your specific areas of weakness so you can focus your study time on the subjects you need most If you want to pass the HESI Admission Assessment Exam or any other admissions assessment exam for health professions that you may encounter there is no better resource than HESI's Admission Assessment Exam Review HESI Hints boxes offer valuable test taking tips as well as rationales suggestions examples and reminders for specific topics Step by step explanations and sample problems in the math section show you how to work through each problem so you understand the steps it takes to complete the equation Sample questions in all other sections prepare you for the questions you will face on the A2 Exam User friendly vocabulary chapter covers more of the medical terminology that you will face on the A2 Exam Easy to read format with consistent section features includes an introduction key terms chapter outline and a bulleted summary to better help you organize your review time and understand the information Full color layout and illustrations visually reinforce key concepts for better understanding NEW 25 question pre

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Teaching Science for Understanding James Joseph Gallagher,2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities Chemistry: The Central Science Theodore L. Brown,H. Eugene LeMay Jr.,Bruce E. Bursten,Catherine Murphy,Patrick Woodward,Steven Langford,Dalius Sagatys,Adrian George,2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet

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