CHEMISTRY STOCHIOMETRY EXAM LEVEL 1 & 2

1. (a) A student carries out a titration to find the concentration of some sulfuric acid.

The student finds that 25.00 cm³ of 0.0880 mol dm⁻³ aqueous sodium hydroxide, NaOH, is neutralised by 17.60 cm³ of dilute sulfuric acid, H₂SO₄.

 $H_2SO_4(aq) + 2NaOH(aq) \rightarrow Na_2SO_4(aq) + 2H_2O(I)$

(i) Calculate the amount, in moles, of NaOH used.

answer =mol

- [7]
- (ii) Determine the amount, in moles, of H₂SO₄ used.

answer =mol

(iii) Calculate the concentration, in mol dm-3, of the sulfuric acid.

answer =mol dm-3

- (b) After carrying out the titration in (a), the student left the resulting solution to crystallise. White crystals were formed, with a formula of Na₂SO₄•x H₂O and a molar mass of 322.1 g mol⁻¹.
 - (i) What term is given to the '-x H₂O' part of the formula?

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