

THEORETICAL YIELD

Example 1: 0.08 g of calcium carbonate reacts with excess hydrochloric acid according to the following balanced equation. Calculate the theoretical yield of carbon dioxide produced at STP (give the volume).



$$m = 0.08 \text{ g}$$

excess

Theoretical yield?

To get from CaCO_3 to CO_2 we need to use a mole ratio... $\text{CaCO}_3 : \text{CO}_2$

This means we FIRST need to calculate moles of CaCO_3

$$n_{\text{CaCO}_3} = \frac{m}{M}$$

$$= \frac{0.08}{100}$$

$$= 0.0008 \text{ mol}$$

This is the amount of CaCO_3 that's used up. We use this in a mole ratio to calculate the theoretical amount of product (CO_2) produced

1:1 ratio means that the amount of CO_2 produced is the same as the CaCO_3 that reacted



$= 0.0008 \text{ mol of CO}_2 \text{ produced}$

This is the theoretical yield of CO_2 in moles. We must convert to volume!

$$n = \frac{V}{V_m}$$

We use this because CO_2 is a gas!

$$0.0008 = \frac{V}{22.4}$$

$$V = 0.01792 \text{ dm}^3$$

Stoichiometry Guided And Study Workbook Answers

Clyde Metz



Stoichiometry Guided And Study Workbook Answers:

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