

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation
 - Identify the given (with units) and what you want to find (with units)
 - Show set up with units. Check sig figs, give final answer with units and label.
- Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



Before $\underline{2} \text{ mol}$ $\underline{0} \text{ mol}$ $\underline{0} \text{ mol}$

Change $\underline{-2 \text{ mol}}$ $\underline{+2 \text{ mol}}$ _____

After $\underline{0} \text{ mol}$ $\underline{2} \text{ mol}$

$$36 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} = 2.0 \text{ mol H}_2\text{O} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} = 2 \text{ mol H}_2 \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} = 4.0 \text{ g H}_2$$

- Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



Before $\underline{9.94} \text{ mol}$ $\underline{\text{xs mol}}$ $\underline{0} \text{ mol}$

Change $\underline{-9.94 \text{ mol}}$ _____ $\underline{+9.94 \text{ mol}}$

After $\underline{0} \text{ mol}$ $\underline{9.94 \text{ mol}}$

$$581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

- You eat 180.0 g of glucose (90 M&Ms). If glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to breathe in to burn the glucose?



Before $\underline{1.0} \text{ mol}$ $\underline{\text{xs mol}}$ $\underline{0} \text{ mol}$ $\underline{0} \text{ mol}$

Change $\underline{-1.0 \text{ mol}}$ _____ $\underline{-1.0 \text{ mol}}$

After $\underline{0} \text{ mol}$ $\underline{\text{xs mol}}$

$$180.0 \text{ g C}_6\text{H}_{12}\text{O}_6 \times \frac{1 \text{ mol C}_6\text{H}_{12}\text{O}_6}{180 \text{ g C}_6\text{H}_{12}\text{O}_6} = 1.00 \text{ mol C}_6\text{H}_{12}\text{O}_6 \times \frac{6 \text{ mol O}_2}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} = 6.00 \text{ mol O}_2$$

$$6.0 \text{ mol O}_2 \times \frac{32.0 \text{ g O}_2}{1 \text{ mol O}_2} = 192.0 \text{ g O}_2$$

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IM Harris



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These bestselling novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of captivating stories waiting to be discovered.

The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts.

The Secret History is a exceptional and thrilling novel that will keep you wondering until the very end. The novel is a warning tale about the dangers of obsession and the power of evil.

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