

### Qualifying Test Two

Directions: Show all work on separate paper and underline your answers.

Phosphorus is a gas that is used when producing fertilizers. The mass of 1 mol of phosphorus is 30.97 g.

1. The volume of phosphorus gas under STP conditions is \_\_\_\_\_.

2. If phosphorus gas is produced under STP conditions, how many moles of phosphorus gas are produced from 125 g of phosphorus?

3. How many grams of phosphorus gas are produced from 125 g of phosphorus?

4. The gas is collected under STP conditions. How many moles of phosphorus gas are produced from 125 g of phosphorus?

5. How many moles of phosphorus gas are produced from 125 g of phosphorus?

6. How many moles of phosphorus gas are produced from 125 g of phosphorus?

How many moles of phosphorus gas are produced from 125 g of phosphorus?

**0.455 mol**

7. How many moles of phosphorus gas are produced from 125 g of phosphorus?

How many moles of phosphorus gas are produced from 125 g of phosphorus?

**125 g**

How many moles of phosphorus gas are produced from 125 g of phosphorus?

**15.6 mol**

8. How many moles of phosphorus gas are produced from 125 g of phosphorus?

**3.33 grams**

How many moles of phosphorus gas are produced from 125 g of phosphorus?

**$1.2 \times 10^3$  mol**

How many moles of phosphorus gas are produced from 125 g of phosphorus?

**4.2 grams**

9. How many moles of phosphorus gas are produced from 125 g of phosphorus?

**0.8273 mol**

How many moles of phosphorus gas are produced from 125 g of phosphorus?

**$2.0 \times 10^2$  grams**

How many moles of phosphorus gas are produced from 125 g of phosphorus?

**0.577 mol**

# Stoichiometry Think Tac Toe Answer Key

**IEEE Antennas and Propagation  
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