

## Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium

### Chapters 17 & 18

**Essential Skills/ State Standards:**

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)  
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)  
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ( $q = m \cdot c \cdot \Delta T$ ) & latent heat of phase change ( $q = m \cdot \Delta H$ ). (CSS: 7d)  
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)  
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d\*, 8c)  
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure (CSS: 9a)  
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

**Vocabulary**

- |   |   |  |  |
|---|---|--|--|
| <ul style="list-style-type: none"> <li>• Condensation</li> <li>• Freezing</li> <li>• Vaporization</li> <li>• Melting</li> <li>• Boiling</li> <li>• Sublimation</li> </ul> | <ul style="list-style-type: none"> <li>• Deposition</li> <li>• Heat</li> <li>• Temperature</li> <li>• Thermochemistry</li> <li>• Exothermic</li> <li>• Endothermic</li> </ul> | <ul style="list-style-type: none"> <li>• Calorimeter</li> <li>• Joule</li> <li>• Specific heat</li> <li>• <math>q = m \cdot c \cdot \Delta T</math></li> <li>• <math>q = m \cdot \Delta H_{\text{vaporization}}</math></li> <li>• <math>q = m \cdot \Delta H_{\text{fusion}}</math></li> </ul> | <ul style="list-style-type: none"> <li>• Heat of reaction</li> <li>• Le Chatelier's principle</li> <li>• Equilibrium</li> <li>• Activation Energy</li> <li>• Catalyst</li> </ul> |
|---|---|--|--|

**Book Assignments:**

Topic	Read	Required Problems
1. Heat, Temperature & Specific Heat	Section 17-1 p. 511-524	Section 17-p. 524 #2, 5 & 6 Chapter 17 Review p. 548 #16 & 17
3. Chemical Equilibrium & Le Chatelier's Principle	Section 12-3 p. 372-375 Section 18-2 p. 562-568	Chapter 18 Review p. 586 #4, 6 & 7
4. Phase Changes Diagrams & Latent Heat of Phase Change	Section 12-3 p. 378-382	Section 12-3 p. 382 #4-8 Chapter 12 Review p. 389 #19, 28, 31 & 37
5. Collision Theory & Activation Energy	Section 17-3 p. 532-534	Section 17-3 p. 537 #2, 3 & 5
6. Reaction Rates & Catalysts	Section 17-4 p. 538-544	Section 17-4 p. 545 #2-4

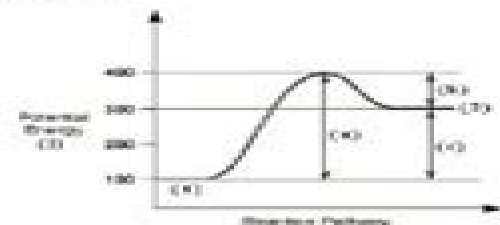
**Study Guide (Use notes, worksheets, and labs in addition to this study guide)**

- According to the collision theory, what 3 things must happen in order for two molecules to react?
- What is the importance of the activation energy in a chemical reaction?
- Using the collision theory, explain if the following would increase/ decrease the rate of reaction:  

$$2 \text{HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$$

a. Increase temp:	b. Increase HCl	c. Increase pressure	d. Remove $\text{MgCl}_2$	e. Add $\text{H}_2$
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- Identify the following terms with their correct number/letter from the graphs:
 

• Reactants-	• Products-
• Activation energy (number/letter & amount in Joules)-	
• Is this reaction exothermic or endothermic?	
- Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do).



# Stoicheometry Study Guide

**James E. Brady**



## **Stoicheometry Study Guide:**

**Chemistry in Everyday Life: A Study Guide** Cybellium, 2024-10-26 Designed for professionals students and enthusiasts alike our comprehensive books empower you to stay ahead in a rapidly evolving digital world Expert Insights Our books provide deep actionable insights that bridge the gap between theory and practical application Up to Date Content Stay current with the latest advancements trends and best practices in IT AI Cybersecurity Business Economics and Science Each guide is regularly updated to reflect the newest developments and challenges Comprehensive Coverage Whether you re a beginner or an advanced learner Cybellium books cover a wide range of topics from foundational principles to specialized knowledge tailored to your level of expertise Become part of a global network of learners and professionals who trust Cybellium to guide their educational journey [www.cybellium.com](http://www.cybellium.com) Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement **Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter *"O" Level Study Guide - Chemistry Quite Easily Done* , *Barron's Science 360: A Complete Study Guide to Chemistry with Online Practice* Barron's Educational Series, Mark Kernion, Joseph A. Mascetta, 2021-09-07 provides a complete guide to the fundamentals of chemistry Page 4 of cover Study Guide to Accompany Calculus for the Management, Life, and Social Sciences Clyde Metz, 1984-01-01 Study Guide to Accompany

Calculus for the Management Life and Social Sciences      **Chemistry, Student Study Guide** John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that's real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You'll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors

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**Practitioners** Ipsita Banerjee, 2025-05-29 This book describes the basic techniques required in learning structural biology Apart from the classical and standard approaches like X ray crystallography and SAXS small angle X ray scattering it has quite an extensive chapter on the most recent approach namely cryo Electron Microscopy The material is partly pedagogical allowing readers to learn the subject including some advanced topics The application of these techniques to a specific example Ribonucleotide Reductase is discussed At the end there is a detailed Mathematical Supplement broken into six sections that develop from basic principles the various mathematical machinery needed to learn the techniques of structural biology It elaborates on the key mathematical concepts mentioned in the various chapters Additionally there is a collection of

Ipsita's papers highlighting her seminal contributions especially in the data processing of cryo EM

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