

Name: _____

Date: _____

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S'mores Stoichiometry Lab



Introduction:

In this experiment, you will explore the principles of stoichiometry by building S'mores, the delicious, chocolate, marshmallow, and graham cracker treats.

Definitions: 'Stoichio' means element and 'metry' means the process of measuring. The mass and quantity relationships among reactants and products in a reaction are found using the process of stoichiometry.

Purpose:

To determine the limiting reactant in the synthesis of S'mores.

Prelab Questions:

1. What does a chemical equation tell us?
2. What do the coefficients in a balanced chemical equation represent?

Materials:

Miniature Chocolate bars
Marshmallows
Bunsen Burner

Graham cracker Square
Paper Towels
Wooden skewers

Reference Data:

Substance	Symbol	Unit mass	Package Mass
Graham Cracker Squares	Gr	8.50 g	408 g
Marshmallow	Mm	7.125 g	285 g
Miniature Chocolate Bar	Ch	12.00 g	348 g

** for the purposes of this lab the UNIT mass and MOLAR mass are the same thing

Procedure:

1. Perform a synthesis reaction (forming ~~one~~ s'more). For this lab, we will consider a s'more to consist of 2 graham cracker squares, 1 marshmallow and 1 miniature chocolate bar.
 - a) Determine the formula for a s'more
 - b) write the balanced chemical equation using the symbols above
 - c) According to the balanced equation complete the following ratios:

_____ Gr = _____ Mm	_____ Gr = _____ S'more
_____ Ch = _____ Mm	_____ Mm = _____ S'more
_____ Ch = _____ Gr	_____ Ch = _____ S'more
2. Calculating the unit mass of the s'more.
 - a) Calculate the mass on ONE s'more.
 - b) Is there a relationship between the mass of a S'more and the masses of the reactants used to make it?
 - c) If so, what is the relationship?
 - d) What law have you studied in this course that might define this relationship?
3. Predicting the limiting reactant:
 - a) Complete the data table below based on the materials provided for your group.

Substance	Quantity received	# of S'mores it COULD make
Graham Cracker Squares		
Marshmallow		
Miniature Chocolate Bar		
 - b) What will the limiting reactant be?
 - c) How many of the other substances will remain
 - d) What are the masses of each excess reactant that will remain once the reaction is complete?

Stoichiometry Lab Key Answers

J Elliott

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Stoichiometry Lab Key Answers:

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underpins the developing field of ecological stoichiometry the study of the balance of chemical elements in ecological interactions This long awaited book brings this field into its own as a unifying force in ecology and evolution Synthesizing a wide range of knowledge Robert Sterner and Jim Elser show how an understanding of the biochemical deployment of elements in organisms from microbes to metazoa provides the key to making sense of both aquatic and terrestrial ecosystems After summarizing the chemistry of elements and their relative abundance in Earth s environment the authors proceed along a line of increasing complexity and scale from molecules to cells individuals populations communities and ecosystems The book examines fundamental chemical constraints on ecological phenomena such as competition herbivory symbiosis energy flow in food webs and organic matter sequestration In accessible prose and with clear mathematical models the authors show how ecological stoichiometry can illuminate diverse fields of study from metabolism to global change Set to be a classic in the field Ecological Stoichiometry is an indispensable resource for researchers instructors and students of ecology evolution physiology and biogeochemistry From the foreword by Peter Vitousek T his book represents a significant milestone in the history of ecology Love it or argue with it and I do both most ecologists will be influenced by the framework developed in this

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Progress in Ecological Stoichiometry Dedmer B. Van de Waal, James J. Elser, Adam C. Martiny, Robert W. Sterner, James B. Cotner, 2018 Ecological stoichiometry concerns the way that the elemental composition of organisms shapes their ecology It deals with the balance or imbalance of elemental ratios and how that affects organism growth nutrient cycling and the interactions with the biotic and abiotic worlds The elemental composition of organisms is a set of constraints through which all the Earth's biogeochemical cycles must pass All organisms consume nutrients and acquire compounds from the environment proportional to their needs Organismal elemental needs are determined in turn by the energy required to live and grow the physical and chemical constraints of their environment and their requirements for relatively large polymeric biomolecules such as RNA DNA lipids and proteins as well as for structural needs including stems bones shells etc These materials together constitute most of the biomass of living organisms Although there may be little variability in elemental ratios of many of these biomolecules changing the proportions of different biomolecules can have important effects on organismal elemental composition Consequently the variation in elemental composition both within and across organisms can be tremendous which has important implications for Earth's biogeochemical cycles It has been over a decade since the publication of Sterner and Elser's book *Ecological Stoichiometry* 2002 In the intervening years hundreds of papers on stoichiometric topics ranging from evolution and regulation of nutrient content in organisms to the role of stoichiometry in populations communities ecosystems and global biogeochemical dynamics have been published Here we present a collection of contributions from the broad scientific community to highlight recent insights in the field of Ecological Stoichiometry

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