

# STOICHIOMETRY: MASS-MASS PROBLEMS

Name \_\_\_\_\_



How many grams of potassium chloride are produced if 25 g of potassium chlorate decompose?

$$25\text{g KClO}_3 \left( \frac{1\text{ mol}}{122.5\text{g}} \right) \left( \frac{2\text{ mol KCl}}{2\text{ mol KClO}_3} \right) \left( \frac{74.5\text{g KCl}}{1\text{ mol KCl}} \right) = 13.2\text{g KCl}$$



How many grams of hydrogen are necessary to react completely with 50.0 g of nitrogen in the above reaction?

$$50\text{g N}_2 \left( \frac{1\text{ mol N}_2}{28\text{g N}_2} \right) \left( \frac{3\text{ mol H}_2}{1\text{ mol N}_2} \right) \left( \frac{2\text{g H}_2}{1\text{ mol H}_2} \right) = 10.7\text{g H}_2$$

3. How many grams of ammonia are produced in the reaction in Problem 2?

$$50\text{g N}_2 \quad 50\text{g} + 10.7 = \boxed{60.7\text{g NH}_3}$$



How many grams of silver chloride are produced from 5.0 g of silver nitrate reacting with an excess of barium chloride?

$$5.0\text{g AgNO}_3 \left( \frac{1\text{ mol}}{170\text{g}} \right) \left( \frac{1\text{ mol AgCl}}{1\text{ mol AgNO}_3} \right) \left( \frac{143\text{g}}{1\text{ mol AgCl}} \right) = 4.21\text{g AgCl}$$

5. How much barium chloride is necessary to react with the silver nitrate in Problem 4?

$$5.0\text{g AgNO}_3 \left( \frac{1\text{ mol}}{170\text{g AgNO}_3} \right) \left( \frac{1\text{ mol BaCl}_2}{2\text{ mol AgNO}_3} \right) \left( \frac{207\text{g}}{1\text{ mol BaCl}_2} \right) = \boxed{5.0\text{g BaCl}_2}$$

9-7

# Stoichiometry Mole Molecule Problems Chemistry If8766

**Michell J. Sienko**



## **Stoichiometry Mole Molecule Problems Chemistry If8766:**

**You Can Do Chemistry** Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018

**Bite-Sized Chemistry Calculations** John Lambert, 2024-07-10 Bite Sized Chemistry Calculations is a series of books on chemistry calculations aimed at helping students overcome the challenges associated with tackling the various types of calculations encountered in different aspects of chemistry focusing on a few topics at a time to facilitate comprehension Written by an experienced chemistry educator each book in the series has been tailored to fully meet the needs of students at all levels especially those taking college level general chemistry courses as well as those following various O level curricula worldwide This part of the series explores the different types of problems and calculations encountered in mass the mole and stoichiometry including the determination of formulae of ionic compounds relative formula masses mass and percent compositions of compounds all aspects of mole calculations empirical and molecular formulae calculations based on chemical equations limiting reagents gas stoichiometry and percent yield The series is packed with many salient features that are meant to facilitate both teaching and learning Some of these include helpful explanations many examples alternative ways to solve problems plenty of practice questions complete answers an appendices With this book you will be well prepared for your exams and boost your performance CONTENTS 1 Writing the Formulae of Ionic Compounds 2 Formula Masses 3 Mass and Percent Compositions 4 The Mole and Mass 5 The Mole and Number of Particles 6 The Mole and Concentration 7 The Mole and Molar Volume 8 Empirical and Molecular Formulae 9 Chemical Equations 10 Calculations Based on Chemical Equations 11 Limiting Reagent 12 Gas Volume Volume Stoichiometry 13 Percent Yield Answers to Practice Problems Appendices Solving Problems in Chemistry Rod O'Connor, Charles Mickey, Alton Hassell, 1977 Stoichiometry and structure Michell J. Sienko, 1964 **Problems and Worked Examples in Chemistry** Albert Holderness, John Lambert, 1978

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**Chemistry Problems** David E. Newton, 2001 This edition includes acid base chemistry and thermochemistry Chemistry Problems is the authoritative resource for practice problems covering all the essentials Includes Atomic structure Stoichiometry Solutions chemistry and Electrochemistry Literally thousands of problems in this compendium build proficiency analytical skills and math skills The text includes a complete answer key and reference to applicable web sites

**Stoichiometry** Michell J. Sienko, 1964 Calculating in Chemistry Cosmas C Mbuji, 2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities Students in higher courses may find it a helpful revision and enhance their clarity The book completely discusses the key topics in basic stoichiometry including the mole concept

reacting quantities and empirical and molecular formulas It begins with the basic concepts and formulas required to convert various quantities to moles or amount of substance This is particularly useful to a beginner Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems The chapter also describes and applies the concept of a limiting reagent Two methods of determining a limiting reagent are explained and illustrated The concept of limiting reagent is extended to reacting volumes of gases with a short cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry *Chemistry Problems* Michell J. Sienko, 1967

**A Stoichiometry Unit** David Callaghan, 2004 **Stoichiometry and Structure** Michell J. Sienko, 1964

*Mathematical Problems for Chemistry Students* Gyorgy Pota, 2011-08-30 *Mathematical Problems for Chemistry Students* has been compiled and written to help chemistry students in their mathematical studies by providing them with mathematical problems really occurring in chemistry b to help practising chemists to activate their applied mathematical skills and c to introduce students and specialists of the chemistry related fields physicists mathematicians biologists etc into the world of the chemical applications Some problems of the collection are mathematical reformulations of those in the standard textbooks of chemistry others were taken from theoretical chemistry journals All major fields of chemistry are covered and each problem is given a solution This problem collection is intended for beginners and users at an intermediate level It can be used as a companion to virtually all textbooks dealing with scientific and engineering mathematics or specifically mathematics for chemists Covers a wide range of applications of the most essential tools in applied mathematics A new approach to a number of classical textbook problems A number of non classical problems are included *Counting Moles* Nigel P. Freestone, 2016-01-04 Students studying chemistry often struggle with the mole *Counting Moles* provides an effective aid to learning by giving clear and confident presentation of the essentials of the mole concept needed by those starting chemistry courses This user friendly self teach e book is split into six chapters which sequentially introduce the mole calculating frame to help solve problems Over 200 fully worked examples are given along with several hundred questions The mole concept is applied to topics such as relative atomic mass and relative formula mass percentage composition empirical

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An Investigation Into Student Difficulties in Advanced Level Chemistry Problems Involving Stoichiometry Brian R. Jolliffe,1983 Stoichiometry and Structure. Freshman Chemistry Problems and how to Solve Them Michell Joseph

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,1964 **Chemistry, Problem-Solving Worktext** James E. Brady,Frederick Senese,2002-06-13 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of  $C_2H_2$  and  $O_2$  to give  $CO_2$  and  $H_2O$  Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet

Formulation and Stoichiometry Emil J. Margolis,2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction **Solving Problems in**

**Chemistry** R. O'Connor, C. Mickey, 1974

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