

## 12

## STOICHIOMETRY

## Practice Problems

In your notebook, solve the following problems.

## SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups (top and bottom), and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate,  $2\text{KClO}_3(\text{s})$ . Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from  $3.8 \times 10^{23}$  molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

## SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(II) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 75.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. Write a balanced equation first.

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of dinitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



# Stoichiometry Section 12 The Arithmetic Of Equations Answers

**M Walker**



## **Stoichiometry Section 12 The Arithmetic Of Equations Answers:**

General College Chemistry Charles William Keenan, Donald C. Kleinfelter, Jesse Hermon Wood, 1980      **Resources in Education** ,1980      *Computers & Electronics* ,1985      **The Software Finder** ,1983      General Chemistry James E. Brady, Gerard E. Humiston, 1986 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and streamlined Changes include more accessible introductory chapters with greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics Each chapter ends with review questions and problems      *Introduction to Chemistry* Amos Turk, 2013-07-15 Introduction to Chemistry is a 26 chapter introductory textbook in general chemistry This book deals first with the atoms and the arithmetic and energetics of their combination into molecules The subsequent chapters consider the nature of the interactions among atoms or the so called chemical bonding This topic is followed by discussions on the nature of intermolecular forces and the states of matter This text further explores the statistics and dynamics of chemistry including the study of equilibrium and kinetics Other chapters cover the aspects of ionic equilibrium acids and bases and galvanic cells The concluding chapters focus on a descriptive study of chemistry such as the representative and transition elements organic and nuclear chemistry metals polymers and biochemistry Teachers and undergraduate chemistry students will find this book of great value

**Chemistry, an Experimental Science** Chemical Education Material Study, 1963      **Maths for Chemistry** Paul Monk, Lindsey J. Munro, 2021 Mathematical skills and concepts lie at the heart of chemistry yet they are the aspect of the subject that many students fear the most Maths for Chemistry recognizes the challenges faced by many students in equipping themselves with the maths skills necessary to gain a full understanding of chemistry Working from foundational principles the book builds the student's confidence by leading them through the subject in a steady progressive way from basic algebra to quantum mathematics Opening with the core mathematics of algebra logarithms and trigonometry the book goes on to cover calculus matrices vectors complex numbers and laboratory mathematics to cover everything that a chemistry student needs With its modular structure the book presents material in short manageable sections to keep the content as accessible and readily digestible as possible Maths for Chemistry is the perfect introduction to the essential mathematical concepts which all chemistry students should master      **Chemistry** Raymond Chang, 1988      *Bowker's Complete Video Directory* ,2000      *Metallurgical & Chemical Engineering* ,1912      *Choice* ,1980      *Electrochemical and Metallurgical Industry* ,1912      **The Computing Teacher** ,1979      **Principles of physical Science** ,1971      Current Index to Journals in Education ,1989      Introduction to General Chemistry Jerome K. Holmes, 1969      **The Merck Report** ,1908      The Complete Home Learning Sourcebook Rebecca Rupp, 1998 Lists all the resources needed to create a balanced curriculum for homeschooling from preschool to high school level      **Iterations** John W. Moore, 1981

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