## Stoichiometry: Multiple Choice

1. The limiting reactant of a chemical reaction is A) the reactant of a chemical process that is not consumed completely B) the product of a chemical process that is consumed completely C) the product of a chemical process that is not consumed completely D) the reactant of a chemical process that is consumed completely E) none of the above ANS: D Consider the following reaction: Ca(OH)<sub>N=0</sub> + Al<sub>2</sub>(SO<sub>4</sub>)<sub>N=0</sub> → CaSO<sub>m0</sub> + Al(OH)<sub>N0</sub> When balanced, the mole ratio for calcium hydroxide reacting with aluminum sulphate is B) 2:1 ANDELL C) 2:3 4D0 1:2: E) 3:11 ANS: E. Sodium phosphate, Na,PO<sub>4</sub>, can be prepared by combining sodium hydroxide and hydrogen phosphate as shown in the following balanced chemical equation: 3NaOH + H.PO, -> Na.PO, + 3H.O If 2.5 mol of H<sub>2</sub>PO<sub>4</sub> are reacted, the number of moles of water that will be produced is A) 1.0 B) 2.5 CO.300 D) 5.0 E3 7.5 ANS: E Potassium chloride and oxygen gas are formed upon heating potassium chlorate. The following: balanced equation shows this decomposition process:  $2KClO_2 \rightarrow 2KCl + 3O_2$ The number of moles of KClO<sub>1</sub> required to produce 5.0 mol of oxygen is A1 1.0 B) 2.0 C0.3.0D0 3.3 EN 5.00 ANS: D Determine the number of grams of CO<sub>2</sub> that can be produced from 48.0 g of the hydrocarbon C<sub>2</sub>H<sub>4</sub>. according to the following equation:  $2C_3H_4 + 9O_5 \rightarrow 6CO_5 + 6H_4O_4$ A) 37.6 g B) 50.2 g: C) 87.8 g D) 151 g E) 188.4 g The limiting reactant in a chemical reaction is the that is A) Reactant, in excess B) Product, left over C) Reactant, used up D) Product, used up

AINS: C

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