Name _____

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- a. Write the balanced chemical equation
- b. Identify the given (with units) and what you want to find (with units)
- c. Show set up with units. Check sig figs, give final answer with units and label.
- Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?

Equation:	2 H ₂ O ₍₀₎ →	2 H _{2(c)} +	1 O2 (a)
Before	2 mol	0 mol	0 mol
Change	-2 mol	+2 mol	
After	0 mol	2 mol	

$$36 \text{ g } H_2Ox \frac{1 \text{ mol } H_2O}{18.0 \text{ g } H_2O} = 2.0 \text{ mol } H_2O \times \frac{2 \text{ mol } H_2}{2 \text{ mol } H_2O} = 2 \text{ mol } H_2 \times \frac{2.0 \text{ g } H_2}{1 \text{ mol } H_2} = 4.0 \text{ g } H_2$$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?

Equation:	2 Na ₍₀₎ +	1 Cl _{2(a)} \rightarrow	2 NaCl _{tet}	
Before	9.94 mol	xs mol	0 mool	
Change	-9.94 mol		+9.94 mol	
After	0 mol		9.94 mod	

$$581 \text{ g NaClx} \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

3. You eat 180.0 g of glucose (90 M&Ms). If glucose, C6H12O6, reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to breathe in to burn the glucose?

Equation:	1 C6H12O6(a) +	6 O _{2(g)} → 6 CO _{2(g)} + 6 H ₂	$O_{(ab)}$
Before	1.0 mod	xs mol 0 mol 0 mol	
Change	-1.0 mol	-1.0 mol	
After	0 mol	xs mol	

$$180.0 \text{ g C}_{a}\text{H}_{12}\text{O}_{a} \times \frac{1 \text{ mol C}_{a}\text{H}_{12}\text{O}_{a}}{180 \text{ g C}_{a}\text{H}_{12}\text{O}_{a}} = 1.00 \text{ mol C}_{a}\text{H}_{12}\text{O}_{a} \times \frac{6 \text{ mol O}_{2}}{1 \text{ mol C}_{a}\text{H}_{12}\text{O}_{a}} = 6.00 \text{ mol O}_{2}$$

$$6.0 \text{ mol } O_2 x \frac{32.0 \text{ g } O_2}{1 \text{ mol } O_2} = 192.0 \text{ g } O_2$$

Stoichiometry Section 2 Answers

John T. Moore, Richard H. Langley

Stoichiometry Section 2 Answers:

Regents Exams and Answers: Chemistry-Physical Setting Revised Edition Barron's Educational Series, Albert Tarendash, 2021-01-05 Barron's Regents Exams and Answers Chemistry provides essential practice for students taking the Chemistry Regents including actual recently administered exams and thorough answer explanations for all questions This book features Eight actual administered Regents Chemistry exams so students can get familiar with the test Thorough explanations for all answers Self analysis charts to help identify strengths and weaknesses Test taking techniques and strategies A detailed outline of all major topics tested on this exam A glossary of important terms to know for test day

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Proceedings of the Third International Symposium on High Temperature Lamp Chemistry Joseph M.
Ranish, Charles W. Struck, 1993

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