

## Study Guide – Chapter 3: Stoichiometry

Test Date: Thursday, 12/18/14

- **Chemical Equations**
  - Writing word equations
    - remember to balance & include the states of matter
  - Balancing chemical equations w/coefficients
  - 5 types of reactions
    - Combination/synthesis, Decomposition, Combustion, Single Replacement, Double Replacement
    - identify what type of reaction is happening from a given equation
    - predict the products when given the reactants or told the type of reaction...
      - $C_xH_y + O_2(g)$  combustion forms products of  $CO_2(g)$  and  $H_2O(g)$
      - Synthesis of metal + nonmetal forms an ionic compound (criss cross charges for formula)
      - Decomposition breaks compound into its elements (watch out for diatomic BrINClHOFs or  $P_4$  /  $S_8$  solids)
- **Formula Weight (unit: amu) = Molar Mass (unit: g/mol)**
  - Calculations of formula weights and/or molar masses
  - % by mass = % composition calculations
- **Determining Empirical & Molecular Formulas**
  - Know the difference between an empirical and a molecular formula
  - Three types of problems
    - Normal problems will give % composition data or mass data to start
    - Combustion Analysis (like the quiz)
    - Hydrate Determination
      - Use the rhyme to help you remember the steps!
- **Stoichiometric Calculations**
  - Dimensional Analysis!!!
  - Avogadro's number (use for converting to or from # of atoms/molecules) → after molecules you can go to specific atom within the compound by looking at the chemical formula (ex: 4 H atoms per 1 molec.  $CH_4$ )
  - use the molar mass to go from mass to moles and/or moles to mass
  - Molar Ratios: use coefficients from a balanced chemical equation to go between any two different substances in a reaction (section 3.6)
- **Limiting Reactant Problems**
  - through calculations, identify which is the limiting reactant and which is the excess reactant
  - calculate the amount of product formed by the limiting reactant (this is the maximum or theoretical yield but you ALWAYS choose the LESSER quantity calculated)
  - determine how much of the excess reactant is used and how much is left over after the reaction is completed
  - calculate % yield
    - calculate theoretical yield and given the actual yield (from lab) of the substance
    - given % yield, you can determine actual yield collected in the lab

# Stoichiometry Study Guide Answer

**CS Toh**



## **Stoichiometry Study Guide Answer:**

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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