

## Study Guide – Chapter 3: Stoichiometry

Test Date: Thursday, 12/18/14

- **Chemical Equations**
  - Writing word equations
    - remember to balance & include the states of matter
  - Balancing chemical equations w/coefficients
  - 5 types of reactions
    - Combination/synthesis, Decomposition, Combustion, Single Replacement, Double Replacement
    - identify what type of reaction is happening from a given equation
    - predict the products when given the reactants or told the type of reaction...
      - $C_xH_y + O_2(g)$  combustion forms products of  $CO_2(g)$  and  $H_2O(g)$
      - Synthesis of metal + nonmetal forms an ionic compound (criss cross charges for formula)
      - Decomposition breaks compound into its elements (watch out for diatomic BrINClHOFs or  $P_4$  /  $S_8$  solids)
- **Formula Weight (unit: amu) = Molar Mass (unit: g/mol)**
  - Calculations of formula weights and/or molar masses
  - % by mass = % composition calculations
- **Determining Empirical & Molecular Formulas**
  - Know the difference between an empirical and a molecular formula
  - Three types of problems
    - Normal problems will give % composition data or mass data to start
    - Combustion Analysis (like the quiz)
    - Hydrate Determination
      - Use the rhyme to help you remember the steps!
- **Stoichiometric Calculations**
  - Dimensional Analysis!!!
  - Avogadro's number (use for converting to or from # of atoms/molecules) → after molecules you can go to specific atom within the compound by looking at the chemical formula (ex: 4 H atoms per 1 molec.  $CH_4$ )
  - use the molar mass to go from mass to moles and/or moles to mass
  - Molar Ratios: use coefficients from a balanced chemical equation to go between any two different substances in a reaction (section 3.6)
- **Limiting Reactant Problems**
  - through calculations, identify which is the limiting reactant and which is the excess reactant
  - calculate the amount of product formed by the limiting reactant (this is the maximum or theoretical yield but you ALWAYS choose the LESSER quantity calculated)
  - determine how much of the excess reactant is used and how much is left over after the reaction is completed
  - calculate % yield
    - calculate theoretical yield and given the actual yield (from lab) of the substance
    - given % yield, you can determine actual yield collected in the lab

# Stoichiometry Study Guide Answer Key

**RJ Alexander**

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

## **Stoichiometry Study Guide Answer Key:**

## Unveiling the Energy of Verbal Art: An Emotional Sojourn through **Stoichiometry Study Guide Answer Key**

In a global inundated with displays and the cacophony of fast connection, the profound power and mental resonance of verbal art usually disappear into obscurity, eclipsed by the regular barrage of sound and distractions. Yet, set within the lyrical pages of **Stoichiometry Study Guide Answer Key**, a interesting function of fictional splendor that pulses with organic emotions, lies an memorable trip waiting to be embarked upon. Composed by way of a virtuoso wordsmith, that mesmerizing opus guides viewers on a psychological odyssey, lightly exposing the latent potential and profound influence embedded within the delicate internet of language. Within the heart-wrenching expanse of this evocative evaluation, we can embark upon an introspective exploration of the book is key themes, dissect its interesting publishing fashion, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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