

Solution Stoichiometry

1. What volume of a 0.35M AgNO_3 is required to completely react with 65 ml of a 0.29M NaCl solution?



Required: V of AgNO_3

Start with the other substance

$$\frac{0.29 \text{ mol NaCl}}{1 \text{ L}} \cdot \frac{0.055 \text{ L NaCl}}{1 \text{ NaCl}} \cdot \frac{1 \text{ mol AgNO}_3}{1 \text{ mol NaCl}}$$

then convert

mol \rightarrow L

$$\cdot \frac{1 \text{ L AgNO}_3}{0.35 \text{ mol AgNO}_3}$$

$$= 0.0377 \text{ L} = 37.7 \text{ mL}$$

2. 37 mL of a 0.29M KI solution is required to completely react with 42 mL of a $\text{Pb}(\text{NO}_3)_2$ solution. What was the original concentration of the lead (II) Nitrate solution?



Given:

V_2 37 mL KI sol
 M_2 0.29M KI sol
 M_1 42 mL $\text{Pb}(\text{NO}_3)_2$

Required:

original concentration of the lead

$$\frac{0.29 \text{ mol KI}}{1 \text{ L}} \cdot \frac{0.037 \text{ L}}{1} \cdot \frac{1 \text{ mol Pb}(\text{NO}_3)_2}{2 \text{ mol KI}} \cdot \frac{1}{0.042 \text{ L}}$$

$$= 0.1277$$

$$2(42) M_2 = 1(0.29)(37)$$

$$M_2 = 10.75 / 84$$

$$M_2 = 0.1277 \text{ M}$$

3. What mass of solid Magnesium Hydroxide can be produced if 15 mL of a 0.43M $\text{Mg}(\text{NO}_3)_2$ solution reacts completely with excess NaOH ?



Required: mass of solid Magnesium Hydroxide if 15 mL of a 0.43M $\text{Mg}(\text{NO}_3)_2$

Given: 15 mL $\text{Mg}(\text{NO}_3)_2$, 0.43M $\text{Mg}(\text{NO}_3)_2$

$$\frac{0.43 \text{ mol Mg}(\text{NO}_3)_2}{1 \text{ L}} \cdot \frac{0.015 \text{ L}}{1} \cdot \frac{1 \text{ mol Mg}(\text{OH})_2}{1 \text{ mol Mg}(\text{NO}_3)_2}$$

$$\cdot \frac{58.32 \text{ g}}{1 \text{ mol Mg}(\text{OH})_2} = 1.65 \text{ g Mg}(\text{OH})_2$$

4. What mass of solid Calcium Phosphate can be produced if 78 mL of a 0.17M Sodium Phosphate solution is mixed with 58 mL of a 0.29M Calcium chloride solution?



Given:

78 mL Na_3PO_4
0.17M Na_3PO_4

58 mL CaCl_2
0.29M CaCl_2

Required:

mass (g) $\text{Ca}_3(\text{PO}_4)_2$

$$\frac{0.17 \text{ mol Na}_3\text{PO}_4}{1 \text{ L}} \cdot \frac{0.078 \text{ L}}{1} \cdot \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{2 \text{ mol Na}_3\text{PO}_4}$$

$$\cdot \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 1.925 \text{ g Ca}_3\text{PO}_4$$

From CaCl_2

$$\frac{0.29 \text{ mol}}{1 \text{ L}} \cdot \frac{0.058 \text{ L}}{1} \cdot \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{3 \text{ mol CaCl}_2} \cdot \frac{310 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol}}$$

$$= 1.9379 \text{ g Ca}_3(\text{PO}_4)_2$$

molar mass:

$$\begin{aligned} \text{Ca} &- 2 \cdot 40.08 \\ \text{P} &- 2 \cdot 30.97 \\ \text{O} &- 8 \cdot 16 \end{aligned} \quad \left. \vphantom{\begin{aligned} \text{Ca} &- 2 \cdot 40.08 \\ \text{P} &- 2 \cdot 30.97 \\ \text{O} &- 8 \cdot 16 \end{aligned}} \right\} 310.18 \text{ g/mol}$$

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Mike Jess



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