



Vocabulary: Stoichiometry



Vocabulary

- Avogadro's number – the number of atoms or molecules in a mole of a substance.
 - Avogadro's number is equal to 6.0221415×10^{23} .
- Balanced equation – an equation that has equal numbers of atoms on each side of the equation.
 - The equation $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ is *unbalanced* because there are two oxygen atoms on the left side and only one on the right.
 - The equation $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ is balanced because there are four hydrogen atoms and two oxygen atoms on each side of the equation.

- Cancel – to cross out, remove, delete, or invalidate something.
 - In an algebraic expression, identical terms that are found in the numerator and denominator can be cancelled, such as the unit “seconds” in this equation:

$$120 \text{ ~~seconds~~} \cdot \frac{1 \text{ minute}}{60 \text{ ~~seconds~~}} = 2 \text{ minutes}$$

- Coefficient – a number that multiplies a term in an equation.
 - In a chemical equation, the coefficients indicate the number of each type of molecule. For example, $6\text{H}_2\text{O}$ means that there are six water molecules.
- Conversion factor – a ratio or fraction equivalent to one that is used to multiply or divide a quantity when converting from one unit to another.
 - For example, the conversion factor for moles of CO_2 to liters of CO_2 is:

$$\frac{1 \text{ mol CO}_2}{22.41 \text{ L CO}_2}$$

This conversion factor is equivalent to one because one mole of carbon dioxide has a volume of 22.41 liters.

- Dimensional analysis – a technique in which comparison of units on both sides of an equation is used to solve problems.
- Formula mass – the mass of one unit of a compound, as measured in unified mass units (u). For example, to find the formula mass of NaCl
 - For example, the formula mass of NaCl is found by adding the average atomic masses of sodium (Na) and chlorine (Cl): $22.99 \text{ u} + 35.45 \text{ u} = 58.44 \text{ u}$.



Stoichiometry Vocabulary Answer

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