



## Worksheet 4

### Chapter 04 | Aqueous Reactions and Solution Stoichiometry

1. What is/are the spectator ion(s) when a  $\text{BaCl}_2$  solution is combined with an  $\text{H}_2\text{SO}_4$  solution?  
☐  $\text{H}^+(\text{aq})$   
☐  $\text{Cl}^-(\text{aq})$   
☐  $\text{SO}_4^{2-}(\text{aq})$ ,  $\text{H}^+(\text{aq})$   
☐  $\text{Ba}^{2+}(\text{aq})$ ,  $\text{H}^+(\text{aq})$   
☐  $\text{Cl}^-(\text{aq})$ ,  $\text{H}^+(\text{aq})$
2. Write a balanced net ionic equation for the reaction, if any, that may occur when the following pair is mixed:  
☐  $2\text{Ag}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s})$   
☐  $2\text{Na}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq}) \rightarrow 2\text{NaNO}_3(\text{s})$   
☐  $2\text{Na}^+(\text{aq}) + 2\text{Ag}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s}) + 2\text{NaNO}_3(\text{s})$   
☐ There are only spectator ions.
3. Write a balanced molecular equation for the reaction of hydrobromic acid with zinc?  
☐  $\text{Zn}(\text{s}) + \text{HBr}(\text{aq}) \rightarrow \text{ZnBr}(\text{aq}) + \text{H}_2(\text{g})$   
☐  $\text{Zn}(\text{s}) + \text{HCl}(\text{aq}) \rightarrow \text{ZnCl}_2(\text{aq}) + \text{H}_2(\text{g})$   
☐  $\text{Zn}(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{ZnCl}_2(\text{aq}) + \text{H}_2(\text{g})$   
☐  $\text{Zn}(\text{s}) + \text{HBr}(\text{aq}) \rightarrow \text{ZnBr}(\text{aq}) + \text{H}^+(\text{aq})$   
☐  $\text{Zn}(\text{s}) + 2\text{HBr}(\text{aq}) \rightarrow \text{ZnBr}_2(\text{aq}) + \text{H}_2(\text{g})$
4. Which of the following generates the most ions in dilute aqueous solution?  
☐  $\text{C}_2\text{H}_5\text{OH}$  covalent compound  $\rightarrow$  no ions  
☐  $\text{CaSO}_4(\text{aq}) \rightarrow \text{Ca}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq})$   
☐  $\text{MgCl}_2(\text{aq}) \rightarrow \text{Mg}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq})$   
☐  $\text{NaCl}(\text{aq}) \rightarrow \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$   
☐  $\text{HBr}(\text{aq}) \rightarrow \text{H}^+(\text{aq}) + \text{Br}^-(\text{aq})$

# Stoichiometry Worksheet Answers Prentice

**N Colangelo**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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