

Periodic Trends: Ionization Energy Worksheet

1. What is the ionization energy of an element?
2. Predict the following trends you expect to see in ionization energy:
 - a) As you move across a period? Why?
 - b) As you move down a group? Why?
3. What factors influence the ease of removing an electron from an element?
4. Choose an orbital where an electron would experience the highest Z_{eff} , effective nuclear charge (least shielded), and the highest first ionization energy. Circle the correct answer.

Na (3s)	Mg (3s)	Al (3p)	P (3p)	S (3p)
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5. Match the following electron configuration with the appropriate ionization energy.

a) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$	i) 1356 kJ/mol
b) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$	ii) 595 kJ/mol
c) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$	iii) 409 kJ/mol
6. Which one of the following has the largest first ionization energy? Circle the correct answer.

Na	K	Li	Cs
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7. The first ionization energies values (kJ/mol) for sequential elements in period 3 are shown below. Assign elements to these values based on the periodic trends.

740	578	786	1012	999
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8. Choose the element with the greatest first ionization energy. Circle the correct answer.

a) Carbon or aluminum	b) Calcium or strontium
c) Helium or lithium	d) Chlorine or argon

Skill Practice 17 Ionization Answers

Victor M. Corman

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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International Conference on Advancements of Medicine and Health Care through Technology; 23 - 26 September 2009 Cluj-Napoca, Romania Simona Vlad,Radu V. Ciupa,Anca I.

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