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Page \_\_\_\_\_



# Moles and Formulas Practice

SCIENCE 10.1.2020

1. How many molecules are in 0.00100 moles of hydrogen gas?

$$\frac{0.00100 \text{ mol H}_2 \times 6.022 \times 10^{23} \text{ molecules/mol}}{1 \text{ mol H}_2} = 6.022 \times 10^{20} \text{ molecules}$$

2. A student needs 11.0 grams of silver (Ag) for an experiment. How many moles is that?

$$\frac{11.0 \text{ g Ag}}{107.87 \text{ g/mol}} = 0.102 \text{ mol Ag}$$

3. How many moles of a metal phosphate,  $\text{M}_3(\text{PO}_4)_2$ , are in 100.0 grams?

$$\frac{100.0 \text{ g M}_3(\text{PO}_4)_2}{353.6 \text{ g/mol}} = 0.283 \text{ mol M}_3(\text{PO}_4)_2$$

4. What is the mass of 0.00100 moles of  $\text{H}_2\text{SO}_4$ ?

$$\frac{0.00100 \text{ mol H}_2\text{SO}_4 \times 98.08 \text{ g/mol}}{1 \text{ mol H}_2\text{SO}_4} = 0.0981 \text{ g H}_2\text{SO}_4$$

5. How many grams of  $\text{H}_2\text{SO}_4$  are in 0.00100 moles?

$$\frac{0.00100 \text{ mol H}_2\text{SO}_4 \times 98.08 \text{ g/mol}}{1 \text{ mol H}_2\text{SO}_4} = 0.0981 \text{ g H}_2\text{SO}_4$$

6. A student does a reaction with 10.0 grams of  $\text{H}_2\text{SO}_4$ . The student finds the mass of the substance in the 10.00g beaker after the reaction is released and is 9.00g. How many grams of  $\text{H}_2\text{SO}_4$  are left?

- a. What is the student's experimental percent yield?

$$\frac{9.00 \text{ g}}{10.0 \text{ g}} \times 100\% = 90.0\%$$

- b. What is the theoretical percent yield?

$$\frac{9.00 \text{ g}}{10.0 \text{ g}} \times 100\% = 90.0\%$$

- c. What was the student's error?



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## **Table of Contents Skill Practice 40 Moles And Formula Practic**

1. Understanding the eBook Skill Practice 40 Moles And Formula Practic
  - The Rise of Digital Reading Skill Practice 40 Moles And Formula Practic
  - Advantages of eBooks Over Traditional Books
2. Identifying Skill Practice 40 Moles And Formula Practic
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Skill Practice 40 Moles And Formula Practic
  - User-Friendly Interface
4. Exploring eBook Recommendations from Skill Practice 40 Moles And Formula Practic
  - Personalized Recommendations
  - Skill Practice 40 Moles And Formula Practic User Reviews and Ratings
  - Skill Practice 40 Moles And Formula Practic and Bestseller Lists
5. Accessing Skill Practice 40 Moles And Formula Practic Free and Paid eBooks

- Skill Practice 40 Moles And Formula Practic Public Domain eBooks
  - Skill Practice 40 Moles And Formula Practic eBook Subscription Services
  - Skill Practice 40 Moles And Formula Practic Budget-Friendly Options
6. Navigating Skill Practice 40 Moles And Formula Practic eBook Formats
- ePub, PDF, MOBI, and More
  - Skill Practice 40 Moles And Formula Practic Compatibility with Devices
  - Skill Practice 40 Moles And Formula Practic Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Skill Practice 40 Moles And Formula Practic
  - Highlighting and Note-Taking Skill Practice 40 Moles And Formula Practic
  - Interactive Elements Skill Practice 40 Moles And Formula Practic
8. Staying Engaged with Skill Practice 40 Moles And Formula Practic
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Skill Practice 40 Moles And Formula Practic
9. Balancing eBooks and Physical Books Skill Practice 40 Moles And Formula Practic
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Skill Practice 40 Moles And Formula Practic
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Skill Practice 40 Moles And Formula Practic
- Setting Reading Goals Skill Practice 40 Moles And Formula Practic
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Skill Practice 40 Moles And Formula Practic
- Fact-Checking eBook Content of Skill Practice 40 Moles And Formula Practic
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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