

Thermochemistry SL IB Paper 2 Practice Questions

1. 22N.2.sl.TZ0.1

Ammonium nitrate, NH_4NO_3 , is used as a high nitrogen fertilizer.

Cold packs contain ammonium nitrate and water separated by a membrane.

(a) The mass of the contents of the cold pack is 25.32 g and its initial temperature is 25.2 °C. Once the contents are mixed, the temperature drops to 0.8 °C.

Calculate the energy, in J, absorbed by the dissolution of ammonium nitrate in water within the cold pack. Assume the specific heat capacity of the solution is $4.18 \text{ J g}^{-1} \text{ K}^{-1}$. Use section 1 of the data booklet.

[1]

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(b) The change in enthalpy when ammonium nitrate dissolves in water is $25.69 \text{ kJ mol}^{-1}$. Determine the mass of ammonium nitrate in the cold pack using your answer obtained in (e)(i) and section 6 of the data booklet.

If you did not obtain an answer in (e)(i), use $3.11 \times 10^3 \text{ J}$, although this is not the correct answer.

[2]

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(c) The absolute uncertainty in mass of the contents of the cold pack is $\pm 0.01 \text{ g}$ and in each temperature reading is $\pm 0.2 \text{ °C}$. Using your answer in (e)(ii), calculate the absolute uncertainty in the mass of ammonium nitrate in the cold pack.

If you did not obtain an answer in (e)(ii), use 6.55 g, although this is not the correct answer.

[3]

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H Kauffman



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