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web feb 22 2022 we use a series of specific numbers called quantum numbers to describe the location of an electron in an associated atom quantum numbers specify the properties of the atomic orbitals and the electrons in those orbitals an electron in an atom or ion has four quantum numbers to describe its state

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web true or false atomic emission spectra can be explained by the wave model of light false or false the neon atoms in a neon sign emit their characteristic color of light as they absorb energy true or false when an atom emits light photons having certain specific energies are being emitted true

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web never true in the quantum mechanical model of the atom the probability of finding an electron within a certain volume of space surrounding the nucleus can be portrayed as a fuzzy cloud always true the fourth principal energy level of an atom contains 32 electrons sometimes true

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web limit an electron's energy to certain values differences between Bohr's model and quantum mechanical model of the atom quantum mechanical model of the atom makes no attempt to describe the electron's path around the nucleus where Bohr's model does atomic orbital