

SOLUBILITY CURVES

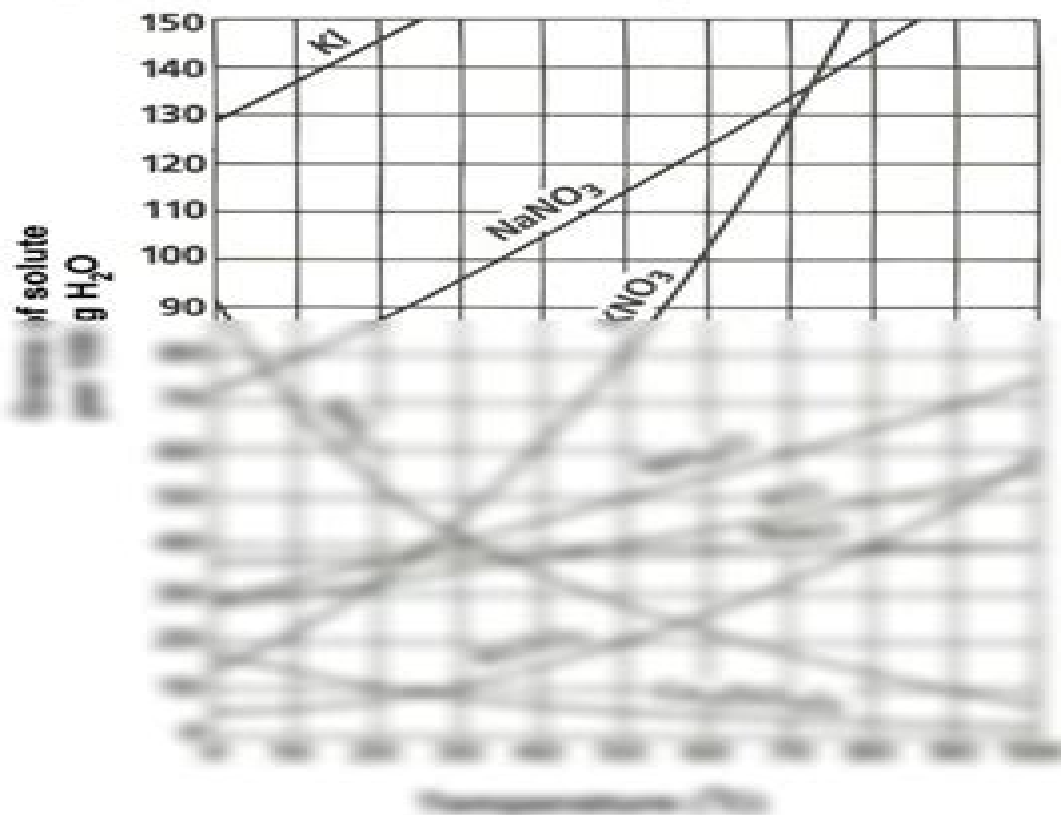
- The lines **on** a solubility curve indicate the concentration of a **saturated** solution – the maximum amount of solute that will dissolve at that specific temperature.
- Values on the graph **beneath** a curve represent **unsaturated** solutions; more solute could be dissolved at that temperature.
- A **supersaturated** solution would have a value **above** the solubility curve at a specific temperature.

Note that since density $\text{H}_2\text{O} = 1.00 \text{ g/mL}$ and since $1\text{mL} = 1\text{cm}^3$:

$$*** 100\text{g H}_2\text{O} = 100 \text{ mL H}_2\text{O} = 100 \text{ cm}^3 \text{ H}_2\text{O}***$$

Example: Find the curve for KClO_3 .

At 30°C approximately 10.0 g of KClO_3 will dissolve in 100g (or 100mL or 100 cm^3) of water. If the temperature is increased to 80°C , approximately 40.0 g of the substance will dissolve in 100g of water.



The most common solubility curve is for sodium chloride, which shows a very slight increase in solubility with temperature. This is because the energy required to break the ionic bonds in the solid is much smaller than the energy released when the ions are hydrated by water molecules.

Most salts are more soluble in water at higher temperatures. However, some salts, such as calcium sulfate, are more soluble in water at lower temperatures. This is because the energy required to break the ionic bonds in the solid is much smaller than the energy released when the ions are hydrated by water molecules.

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