

# QUIZZZ

AP Chemistry Unit 8 Review

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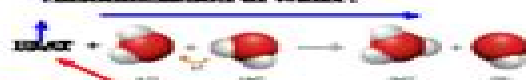
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## Unit 8

### Acids & Bases

#### 8.1 Introduction to Acids & Bases

##### Autoionization of Water:



This is an endothermic process- increasing the heat will shift the reaction to produce more  $\text{H}_3\text{O}^+$  and  $\text{OH}^-$ .  $K_w$  increases (more products), both pH and pOH will decrease but pure H<sub>2</sub>O will still be neutral.

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$$

Only digits to the right of the decimal are significant in pH and pOH values. See

$$\text{pH} = 3.45$$

is sig. figs

In case of acidic solution  $> [\text{H}^+] > [\text{OH}^-]$

$$\text{pH} < 7.00 \text{ at } 25^\circ\text{C}$$

In case of neutral solution  $> [\text{H}^+] = [\text{OH}^-]$

$$\text{pH} = 7.00 \text{ at } 25^\circ\text{C}$$

In case of basic solution  $> [\text{H}^+] < [\text{OH}^-]$

$$\text{pH} > 7.00 \text{ at } 25^\circ\text{C}$$

##### ION PRODUCT CONSTANT:

$$K_w = [\text{H}^+][\text{OH}^-] = 1.0 \times 10^{-14} \text{ at } 25^\circ\text{C}$$

$$= K_a \times K_b$$

$$\text{pH} = -\log[\text{H}^+], \text{pOH} = -\log[\text{OH}^-]$$

$$14 = \text{pH} + \text{pOH}$$

Also known:

$$[\text{H}^+] = 10^{-\text{pH}} \text{ and } [\text{OH}^-] = 10^{-\text{pOH}}$$

3. The value of  $K_w$  for water at  $0^\circ\text{C}$  is  $1 \times 10^{-15}$ . What is the pOH of water at  $0^\circ\text{C}$ ?

a) 7.5

b) 8.0

c) 7.0

d) 6.5

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**J. S. Brimacombe**



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