

Section-D CHEMISTRY

- What happens when dilute sulphuric acid is added to Zn granules?
 - No reaction takes place
 - Chlorine gas and Zinc hydroxide are produced
 - Zinc salt and water are produced
 - Hydrogen gas and Zinc chloride are produced
- Which of the following are characteristics of isotopes of an element?
 - Isotopes of an element show same physical properties
 - Isotopes of an element have same atomic masses
 - Isotopes of an element have same chemical properties
 - Isotopes of an element have same atomic number
- Which of the following hydroxides are most basic?
 - $\text{Ba}(\text{OH})_2$
 - $\text{Ca}(\text{OH})_2$
 - $\text{Ba}(\text{OH})_2$
 - $\text{Mg}(\text{OH})_2$
- The longest and the shortest periods are
 - 1 & 7
 - 3 & 6
 - 1 & 6
 - 4 & 5
- Acids react with bases to form salt and water. This reaction is known as
 - reduction
 - Decomposition
 - Neutralization
 - Combination
- A solution of a substance 'X' is used for white washing. Name the substance 'X' and write its formula.
 - Slime, CaCO_3
 - Slime stone, CaCO_3
 - Calcium carbonate, CaCO_3
 - Calcium oxide, CaO
- How is the concentration of hydronium ions (H_3O^+) affected when a solution of an acid is diluted?
 - Increases
 - Decreases
 - Increases
 - Remains the same

- When a gold foil is bombarded by a beam of α particles, only a few of them get deflected whereas most go straight undeflected. This is because
 - The volume of nucleus is much smaller than that of the atom
 - The force of repulsion acting on the fast moving α particles is very small
 - The force of attraction exerted on α particles by the electrons is insufficient
 - The neutrons have no effect on α particles
- Which of the following does not have 2 valence electrons?
 - He
 - Cl^-
 - Ar
 - Ne
- Molecular weight of $\text{CaCO}_3 \cdot \text{H}_2\text{O}$ is equal to
 - $100.5 + 18 + 16$
 - 100.5
 - 100.5 + 18
 - 100.5×18
- Among the following select the metal found free in nature
 - Au
 - Na
 - Mg
 - Cu
- Which of the following represents cyclohexane?
 - C_6H_{12}
 - C_6H_6
 - C_6H_8
 - C_6H_{10}
- An element X has valency equal to 3. What will be its formula with carbonate ions?
 - $\text{X}(\text{CO}_3)_3$
 - X_2CO_3
 - X_2CO_3
 - $\text{X}_2(\text{CO}_3)_3$
- Na^+ has 11 neutrons and 10 electrons. Which of the following statements is correct?
 - Na^+ has atomic number 10 and mass number 21
 - Na^+ has atomic number 11 and mass number 21
 - Na^+ has atomic number 10 and mass number 22
 - Na^+ has atomic number 11 and mass number 22
- The number of covalent bonds in C_6H_{12} is
 - 18
 - 16
 - 15
 - 17

Sample Paper For Grand Dakshana Selection Test

Didier Musso



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