

Potentially useful data:

Avogadro's constant: $6.022 \times 10^{23} \text{ mol}^{-1}$

1. What is the correct noble gas (shorthand) notation for the electron configuration given below?



- a) [Ar] $5s^2$
 - b) [Ar] $3s^2 3p^6 3d^{10} 4p^6 5s^2$
 - c) [Ca] $3d^{10} 4p^6 5s^2$
 - d) [Kr] $5s^2$
 - e) [Rb] $5s^1$
2. Why does an electron found in a $2s$ -orbital have a lower energy than an electron found in a $2p$ -orbital?
- a) Electrons in the $2s$ -subshell are shielded by electrons in the $2p$ -subshell.
 - b) The shape of the orbital ultimately determines the energy of the electrons.
 - c) The larger number of electrons found in the $2p$ -orbital leads to greater repulsion.
 - d) Electrons in the $2s$ -orbital are at a smaller average relative distance from the nucleus.
 - e) The first energy level does not contain p -orbitals.
3. Which of the following does **NOT** influence the effective nuclear charge experienced by a $3d$ -electron in a titanium atom?
- a) the number of protons in the nucleus of the titanium atom
 - b) the distance between the $3d$ electron and the nucleus
 - c) the shielding of the $3d$ electron by the $4s$ -subshell
 - d) the shielding of the $3d$ electron by the $4f$ -subshell
 - e) All of the above-mentioned variables influence the indicated effective nuclear charge.
4. Arrange the following elements in order of **INCREASING** (that is, from smallest to largest) atomic radius.
- a) K < Na < P < Cl < Si
 - b) Si < Na < P < Cl < K
 - c) Si < Cl < P < Na < K
 - d) Si < Na < K < P < Cl
 - e) Cl < P < Si < Na < K

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**Daniel W Foster Professor of Medical
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Table of Contents Sample Questions On Chemistry 101

1. Understanding the eBook Sample Questions On Chemistry 101
 - The Rise of Digital Reading Sample Questions On Chemistry 101
 - Advantages of eBooks Over Traditional Books
2. Identifying Sample Questions On Chemistry 101
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Sample Questions On Chemistry 101
 - User-Friendly Interface
4. Exploring eBook Recommendations from Sample Questions On Chemistry 101
 - Personalized Recommendations
 - Sample Questions On Chemistry 101 User Reviews and Ratings
 - Sample Questions On Chemistry 101 and Bestseller Lists
5. Accessing Sample Questions On Chemistry 101 Free and Paid eBooks
 - Sample Questions On Chemistry 101 Public Domain eBooks
 - Sample Questions On Chemistry 101 eBook Subscription Services
 - Sample Questions On Chemistry 101 Budget-Friendly Options

6. Navigating Sample Questions On Chemistry 101 eBook Formats
 - ePub, PDF, MOBI, and More
 - Sample Questions On Chemistry 101 Compatibility with Devices
 - Sample Questions On Chemistry 101 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Sample Questions On Chemistry 101
 - Highlighting and Note-Taking Sample Questions On Chemistry 101
 - Interactive Elements Sample Questions On Chemistry 101
8. Staying Engaged with Sample Questions On Chemistry 101
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Sample Questions On Chemistry 101
9. Balancing eBooks and Physical Books Sample Questions On Chemistry 101
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Sample Questions On Chemistry 101
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Sample Questions On Chemistry 101
 - Setting Reading Goals Sample Questions On Chemistry 101
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Sample Questions On Chemistry 101
 - Fact-Checking eBook Content of Sample Questions On Chemistry 101
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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