

Percent Yield Lab

Purpose:

To predict the amount of product generated from a double displacement reaction

Materials:

- Sodium Bicarbonate
- Balance
- Weigh boat
- 500 mL Erlenmeyer flask
- Distilled water
- 25 mL graduated cylinder
- Stir Rod
- Acetic Acid
- 100 mL graduated cylinder
- Hot plate

Safety:

Safety goggles and aprons are to be worn at all times

Pre-Lab Calculation:

1. Calculate how many grams of sodium bicarbonate (NaHCO_3) will need to be measured out to obtain 0.025 moles (all work must be shown here)

Procedure:

1. Measure out _____ of sodium bicarbonate and record this amount in your data table
2. Take the mass of a clean 500 mL Erlenmeyer flask
3. Pour the sodium bicarbonate in the Erlenmeyer flask along with a measured out 15 mL of distilled water. Stir until the solute is mostly dissolved.
4. Measure out 75 mL of acetic acid and SLOWLY add it in small quantities to the flask containing the sodium bicarbonate solution. Wait for the bubbling to subside between additions so the reaction does not overflow (overflow will affect your grade)
5. Once all acetic acid has been added, stir for two minutes
6. When the solution is completely calm, move the flask to a hot plate and heat it to boiling. Be careful that the flask does not overflow!
7. When all liquid has boiled away, remove the flask from the hotplate and allow it to cool to room temperature.
8. When the flask has completely cooled, take the mass of the salt and flask. The salt in the flask is sodium acetate.

Percentage Yield Lab Grade 1

RS Peters



Percentage Yield Lab Grade 1:

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