# CHAPTER 17 STUDY GUIDE

# Chemical Equilibrium

### Section 17.1 Equilibrium: A State of Dynamic Balance

In your textbook, read about chemical equilibrium.

Complete each statement.

- When a reaction results in almost complete conversion of reactants to products, chemists say the reaction goes to
- 2. A reaction that can occur in both the forward and the reverse directions is called a(n)

 is a state in which the forward and reverse reactions balance each other because they take place at equal rates.

- At equilibrium, the concentrations of reactants and products are \_\_\_\_\_\_\_ but that does not mean that the amounts or concentrations are \_\_\_\_\_\_.
- 5. Equilibrium is a state of \_\_\_\_\_\_, not one of \_\_\_\_\_\_

In your textbook, read about equilibrium expressions and constants.

For each statement below, write true or false.

- The law of chemical equilibrium states that at a given pressure, a chemical system may reach a state in which a particular ratio of reactant to product concentrations has a constant value.
  - 7. The equation  $H_2(g) + I_2(g) \leftrightarrow 2HI(g)$  is an example of a homogeneous equilibrium.
- If an equilibrium constant has a value less than one, the reactants are favored at equilibrium.
- 9. The value for K<sub>eq</sub> is constant only at a specific volume.
- If the equilibrium constant for a reaction at 300 K is 49.7, the concentration of the reactants will be greater than the concentration of the products.
- A heterogeneous equilibrium means that reactants and products are present in more than one state.
  - 12. The product of the forward chemical reaction is HI, for the equilibrium expression:

$$K_{eq} = \frac{[HI]^2}{[H_2][I_2]}$$

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**Susan McMurry** 

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