

Percent Composition Practice 1.8 -- Answer Key

Objective: - Calculate the percent composition of a compound.

Directions: Determine the percent composition of each of the compounds below. (Show all work!)

1. HCl

$$\text{H: } 1.00\text{g} \times 1 = 1.01\text{g}$$

$$\left(\frac{1.01\text{g}}{35.45\text{g}} \right) \times 100\% = \boxed{2.73\%}$$

$$\text{Cl: } 35.45\text{g} \times 1 = 35.45\text{g}$$

$$\left(\frac{35.45\text{g}}{35.45\text{g}} \right) \times 100\% = \boxed{97.27\%}$$

2. KMnO_4

$$\text{K: } 39.10\text{g} \times 1 = 39.10\text{g}$$

$$\left(\frac{39.10\text{g}}{158.04\text{g}} \right) \times 100\% = \boxed{24.74\%}$$

$$\text{Mn: } 54.94\text{g} \times 1 = 54.94\text{g}$$

$$\left(\frac{54.94\text{g}}{158.04\text{g}} \right) \times 100\% = \boxed{34.76\%}$$

$$\text{O: } 16.00\text{g} \times 4 = 64.00\text{g}$$

$$\left(\frac{64.00\text{g}}{158.04\text{g}} \right) \times 100\% = \boxed{40.50\%}$$

3. $\text{Mg(NO}_3)_2$

$$\text{Mg: } 24.31\text{g} \times 1 = 24.31\text{g}$$

$$\left(\frac{24.31\text{g}}{148.31\text{g}} \right) \times 100\% = \boxed{16.39\%}$$

$$\text{N: } 14.01\text{g} \times 2 = 28.02\text{g}$$

$$\left(\frac{28.02\text{g}}{148.31\text{g}} \right) \times 100\% = \boxed{18.90\%}$$

$$\text{O: } 16.00\text{g} \times 6 = 96.00\text{g}$$

$$\left(\frac{96.00\text{g}}{148.31\text{g}} \right) \times 100\% = \boxed{64.72\%}$$

4. $(\text{NH}_4)_3\text{PO}_4$

$$\text{N: } 14.01\text{g} \times 3 = 42.03\text{g}$$

$$\left(\frac{42.03\text{g}}{149.12\text{g}} \right) \times 100\% = \boxed{28.19\%}$$

$$\text{H: } 1.01\text{g} \times 12 = 12.12\text{g}$$

$$\left(\frac{12.12\text{g}}{149.12\text{g}} \right) \times 100\% = \boxed{8.13\%}$$

$$\text{P: } 30.97\text{g} \times 1 = 30.97\text{g}$$

$$\left(\frac{30.97\text{g}}{149.12\text{g}} \right) \times 100\% = \boxed{20.77\%}$$

$$\text{O: } 16.00\text{g} \times 4 = 64.00\text{g}$$

$$\left(\frac{64.00\text{g}}{149.12\text{g}} \right) \times 100\% = \boxed{42.92\%}$$

Percent Composition Answers

Chris Hren, Peter J. Mikulecky



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who need to have a better understanding of chemistry will benefit from this accessible book. It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features. Each section focuses on the development and assessment of one or two specific objectives. Within each section a specific objective is included, an anticipatory set to orient the reader, content discussion from established authors and guided practice problems for relevant objectives. These features are followed by a set of independent practice problems. The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics. Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance. **Fundamentals of Environmental Chemistry, Third Edition** Stanley E.

Manahan, 2011-03-05 Written by an expert using the same approach that made the previous two editions so successful. **Fundamentals of Environmental Chemistry Third Edition** expands the scope of book to include the strongly emerging areas broadly described as sustainability science and technology including green chemistry and industrial ecology. The new edition includes increased emphasis on the applied aspects of environmental chemistry. Hot topics such as global warming and biomass energy. Integration of green chemistry and sustainability concepts throughout the text. More and updated questions and answers including some that require Internet research. Lecturers Pack on CD ROM with solutions manual. PowerPoint presentations and chapter figures available upon qualifying course adoptions. The book provides a basic course in chemical science including the fundamentals of organic chemistry and biochemistry. The author uses real life examples from environmental chemistry, green chemistry and related areas while maintaining brevity and simplicity in his explanation of concepts. Building on this foundation the book covers environmental chemistry broadly defined to include sustainability aspects, green chemistry, industrial ecology and related areas. These chapters are organized around the five environmental spheres: the hydrosphere, atmosphere, geosphere, biosphere and the anthrosphere. The last two chapters discuss analytical chemistry and its relevance to environmental chemistry. Manahan's clear, concise and readable style makes the information accessible regardless of the reader's level of chemistry knowledge. He demystifies the material for those who need the basics of chemical science for their trade, profession or study curriculum as well as for readers who want to have an understanding of the fundamentals of sustainable chemistry in its crucial role in maintaining a livable planet. **Fundamentals of**

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General Chemistry Workbook Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

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Table of Contents Percent Composition Answers

1. Understanding the eBook Percent Composition Answers
 - The Rise of Digital Reading Percent Composition Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Percent Composition Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Percent Composition Answers
 - Personalized Recommendations
 - Percent Composition Answers User Reviews and Ratings
 - Percent Composition Answers and Bestseller Lists
5. Accessing Percent Composition Answers Free and Paid eBooks
 - Percent Composition Answers Public Domain eBooks
 - Percent Composition Answers eBook Subscription Services
 - Percent Composition Answers Budget-Friendly Options
6. Navigating Percent Composition Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Percent Composition Answers Compatibility with Devices
 - Percent Composition Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Percent Composition Answers
 - Highlighting and Note-Taking Percent Composition Answers
 - Interactive Elements Percent Composition Answers
8. Staying Engaged with Percent Composition Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Percent Composition Answers
- 9. Balancing eBooks and Physical Books Percent Composition Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Percent Composition Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Percent Composition Answers
 - Setting Reading Goals Percent Composition Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Percent Composition Answers
 - Fact-Checking eBook Content of Percent Composition Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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