

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - (a) according to increasing atomic mass
 - (b) according to Mendeleev's original design
 - (c) according to increasing atomic number
 - (d) based on how they were discovered
2. _____ Mendeleev stated that certain elements have similar chemical properties
 - (a) elements appeared in regular intervals and subsequent elements were arranged based on that spacing
 - (b) elements
 - (c) atomic numbers
 - (d) mass ratios
3. _____ The modern periodic law states that
 - (a) all new elements with the same properties are listed in the same group
 - (b) the physical and chemical properties of elements are functions of atomic number
 - (c) elements exhibit properties of both periods and groups
 - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. _____ The discovery of the noble gases changed Mendeleev's periodic table
 - (a) adding a row
 - (b) adding a new period
 - (c) adding a new column
 - (d) adding a new group
5. _____ The noble gases that occupy the group 8 in Mendeleev's
 - (a) noble gases
 - (b) noble gases
 - (c) noble gases
 - (d) noble gases
6. _____ Likewise, the first column in Group 1 has an atomic number of 3. The second element in this group has an atomic number of
 - (a) 4
 - (b) 5
 - (c) 6
 - (d) 7
7. How many elements are in Group 1?
 - (a) 1
 - (b) 2
 - (c) 3
 - (d) 4
8. How many elements are in Group 2?
 - (a) 1
 - (b) 2
 - (c) 3
 - (d) 4
9. What is the atomic number of the element in Group 1, Period 2?
 - (a) 3
 - (b) 4
 - (c) 5
 - (d) 6

Periodic Law Chapter 5 Review Answers

HESI



Periodic Law Chapter 5 Review Answers:

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