

# Annotated periodic table of the elements

1	H 1.01 Hydrogen																		13	14	15	16	17	18
2	3	4	Most common oxidation states, electronegativity (Pauli-Pauling)										5	6	7	8	9	10						
11	12	13											19	20	21	22	23	24						
37	38	39											55	56	57	58	59	60						
87	88	89											119	120	121	122	123	124						
117	118	119											157	158	159	160	161	162						

Group name\*

1 Hydrogen & the alkali metals

2 Alkaline earth metals

13 Coinage metals

14 Volatile metals

15 Isogagers

14 Crystallogens

15 Pnictogens

16 Chalcogens

17 Halogens

18 Helium group

\*Groups 1 to 10 are named after their first members i.e. Group 1 is the Scandium Group

1 Scandium

2 Titanium

3 Vanadium

4 Chromium

5 Manganese

6 Iron

7 Cobalt

8 Nickel

9 Copper

10 Zinc

11 Gallium

12 Germanium

13 Arsenic

14 Selenium

15 Tellurium

16 Iodine

17 Xenon

18 Astatine

19 Francium

20 Radium

21 Actinium

22 Thorium

23 Protactinium

24 Uranium

25 Neptunium

26 Plutonium

27 Americium

28 Curium

29 Berkelium

30 Californium

31 Einsteinium

32 Fermium

33 Mendelevium

34 Nobelium

35 Lawrencium

36 Rutherfordium

37 Dubnium

38 Seaborgium

39 Bohrium

40 Hassium

41 Meitnerium

42 Darmstadtium

43 Roentgenium

44 Copernicium

45 Dubnium

46 Seaborgium

47 Bohrium

48 Hassium

49 Meitnerium

50 Darmstadtium

51 Roentgenium

52 Copernicium

53 Dubnium

54 Seaborgium

55 Bohrium

56 Hassium

57 Meitnerium

58 Darmstadtium

59 Roentgenium

60 Copernicium

61 Dubnium

62 Seaborgium

63 Bohrium

64 Hassium

65 Meitnerium

66 Darmstadtium

67 Roentgenium

68 Copernicium

69 Dubnium

70 Seaborgium

71 Bohrium

72 Hassium

73 Meitnerium

74 Darmstadtium

75 Roentgenium

76 Copernicium

77 Dubnium

78 Seaborgium

79 Bohrium

80 Hassium

81 Meitnerium

82 Darmstadtium

83 Roentgenium

84 Copernicium

85 Dubnium

86 Seaborgium

87 Bohrium

88 Hassium

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93 Dubnium

94 Seaborgium

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1000 Hassium

- Group names\*
- Hydrogen & the alkali metals
  - Alkaline earth metals
  - Coinage metals
  - Volatile metals
  - Isotopes
  - Chalcogens
  - Postcogens
  - Chalcogens
  - Halogens
  - Helium group

\*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

Main group  
oxidation states  
electronegativity  
(Alred-Pauling)

Atomic number  
and symbol  
Atomic weight  
and name

## CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Active
Transition	Intermediate
Post-transition	Metals
Metals (Ru-Pt, Os-Au)	Metals

57 La 138.91 Lanthanum	58 Ce 140.12 Cerium	59 Pr 140.91 Praseodymium	60 Nd 144.24 Neodymium	61 Pm 144.91 Promethium	62 Sm 150.36 Samarium	63 Eu 151.96 Europium	64 Gd 157.25 Gadolinium	65 Tb 158.93 Terbium	66 Dy 162.50 Dysprosium	67 Ho 164.93 Holmium	68 Er 167.26 Erbium	69 Tm 168.93 Thulium	70 Yb 173.05 Ytterbium
89 Ac 227.03 Actinium	90 Th 232.04 Thorium	91 Pa 231.04 Protactinium	92 U 238.03 Uranium	93 Np 237.05 Neptunium	94 Pu 244.06 Plutonium	95 Am 243.06 Americium	96 Cm 247.07 Curium	97 Bk 247.07 Berkelium	98 Cf 251.08 Californium	99 Es 252.08 Einsteinium	100 Fm 257.10 Fermium	101 Md 258.10 Mendelevium	102 No 259.10 Nobelium

Radioactive

Approximate  
likely to be either  
a metal or a  
post-transition  
metal

## Active metals

- Pre-transition**  
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).
- Rare earths**  
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphoteric.
- Actinides**  
Soft, dense and reactive metals. Those from Th onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

## Transition metals

- Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Ir onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.
- Noble metals**  
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very easily electroplated; insoluble oxides.
- Frontier metals**  
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphoteric, and anionic species such as zincates, stannates, and bismuthates.

## Metals

- Metalloids**  
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.
- Intermediate nonmetals**  
H and N are colourless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H<sub>2</sub>O, CO, NO, N<sub>2</sub>O).
- Group 1**  
The alkali metals Li, Na, K, Rb, Cs, and Fr are soft, silvery-white, and highly reactive. They are the most reactive of the elements. They are found in nature as compounds, not as free elements. They are the most reactive of the elements. They are found in nature as compounds, not as free elements.

## Nonmetals

- Chalcogens**  
O, S, Se, and Te are, respectively, colourless, pale yellow, and pale yellow-green solids. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.
- Noble gases**  
Colourless, odourless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF<sub>6</sub>, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

# Periodic Table Guide Answer Sheet

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