

Worksheet 1: Periodic properties and variation of properties.

1. The following table represents the first three periods of the modern periodic table. Study the table and answer the questions that follow:

1 H 1.008							2 He 4.003
3 Li 6.941	4 Be 9.012	5 B 10.81	6 C 12.01	7 N 14.01	8 O 16.00	9 F 18.998	10 Ne 20.18
11 Na 22.99	12 Mg 24.31	13 Al 26.98	14 Si 28.09	15 P 30.97	16 S 32.07	17 Cl 35.45	18 Ar 39.95
19 K 39.10	20 Ca 40.08						

- (i) Write the formula of the sulphate of the element with atomic number 13.
- (ii) What type of bonding will be present in the oxide of the element with atomic number 1?
- (iii) Name the elements which have the highest ionization potential.
- (iv) How many electrons are present in the valence shell of the element with the atomic number 18?
- (v) What is the name given to the energy released when an atom in its isolated gaseous state accepts an electron to form an anion?
- (vi) What is the electronic configuration of the element in the third period which gains one electron to change into an anion?

2. Fill in the blanks:

- i) The atomic size _____ as we move from left to right across the period, because the _____ increases but the _____ remains the same.
- (ii) Choose the word or phrase from the brackets which correctly completes each of the following statements:
- (iii) The element below sodium in the same group would be expected to have a _____

Periodic Variation Of Properties Answer Key

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