# Examen de révision de chimie des solutions

#### Exercice 1

Compléter le tableau suivant :

Solution Aqueuse	PKa	Co en mol/l	[H3O+] en mol/I	[HO-] en mol/l	[Acide] en mol/l	[Base] en mol/l	рН	Coefficient d'ionisation	Type de Solution
Acide		0.001					3		
HCN	9.4	Décimolaire	le ji						
NaOH		Molaire							
CH3NH2	10.6	Centimolaire	r i	-1	1-7				

#### Exercice 2

On considère une solution S1 d'acide faible HBO2 de concentration C1=10<sup>-3</sup>M.

- Calculer le pH de cette solution (Démonstration)
- 2. Calculer le coefficient de dissociation
- A 1 litre de la solution de HBO<sub>2</sub> précédente, on ajoute 1 litre d'une solution d'acide fort HCl de concentration C<sub>2</sub>=10<sup>-2</sup>M.
- 3. Donner l'expression de H3O+ de la solution obtenue.
- 4. Calculer le pH de cette solution

A litre de la solution S1 de HBO2, on ajoute sans variation, 10g du sel NaBO2, totalement soluble dans l'eau.

- Ecrire les réactions qui ont lien dans la solution.
- Quelle est la nature de la solution obtenue.
- Calculer le Ph.

pKa (HBO2/BO2-) = 9.16; Masse molaire de NaBO2 est 66 g/mol

# Exercice 3

1. Déterminer la solubilité en mol/l des composés suivants :

Sel	CaSO4	AgBr	PbSO4	AgCl	
Produit de solubilité (Ks)	2.0 10-5	7.7 10-3	2.0 10-8	1.6 10-10	

Y aura-t-il précipitation :

- Si On mélange 100 ml de Ca/Cl<sub>2</sub> 0.02 M avec 100 ml de Na<sub>2</sub>SO<sub>4</sub> 4.10<sup>-4</sup> M.
- Si on mélange 250 ml d'une solution de Pb(NO3)2 1.6 10<sup>-3</sup> M avec 750 ml d'une solution de Na2SO4 2.4 10<sup>-3</sup> M.
- 4. On ajoute une solution de nitrate d'argent concentrée goutte à goutte à une solution aqueuse dont la concentration en ions Cl' est de 0.1M et la concentration en ions Br- est de 10<sup>-4</sup> M. Quelle est le sel qui précipité le premier AgCl ou AgBr ??

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