

## Examen de révision de chimie des solutions

### Exercice 1

Compléter le tableau suivant :

Solution Aqueuse	PKa	Co en mol/l	[H <sub>3</sub> O <sup>+</sup> ] en mol/l	[HO <sup>-</sup> ] en mol/l	[Acide] en mol/l	[Base] en mol/l	pH	Coefficient d'ionisation	Type de Solution
Acide		0,001					3		
HCN	9,4	Décimolaire							
NaOH		Molaire							
CH <sub>3</sub> NH <sub>2</sub>	10,6	Centimolaire							

### Exercice 2

On considère une solution S1 d'acide faible HBO<sub>2</sub> de concentration C<sub>1</sub>=10<sup>-2</sup>M.

1. Calculer le pH de cette solution (Démonstration)
2. Calculer le coefficient de dissociation

A 1 litre de la solution de HBO<sub>2</sub> précédente, on ajoute 1 litre d'une solution d'acide fort HCl de concentration C<sub>2</sub>=10<sup>-2</sup>M.

3. Donner l'expression de H<sub>3</sub>O<sup>+</sup> de la solution obtenue.
4. Calculer le pH de cette solution

A litre de la solution S1 de HBO<sub>2</sub>, on ajoute sans variation, 10g du sel NaBO<sub>2</sub>, totalement soluble dans l'eau.

5. Ecrire les réactions qui ont lieu dans la solution.
6. Quelle est la nature de la solution obtenue.
7. Calculer le Ph.

pKa (HBO<sub>2</sub>/BO<sub>2</sub><sup>-</sup>) = 9,16 ; Masse molaire de NaBO<sub>2</sub> est 66 g/mol

### Exercice 3

1. Déterminer la solubilité en mol/l des composés suivants :

Sel	CaSO <sub>4</sub>	AgBr	PbSO <sub>4</sub>	AgCl
Produit de solubilité (Ks)	2,0 10 <sup>-6</sup>	7,7 10 <sup>-13</sup>	2,0 10 <sup>-8</sup>	1,6 10 <sup>-10</sup>

Y aura-t-il précipitation :

2. Si On mélange 100 ml de CaCl<sub>2</sub> 0,02 M avec 100 ml de Na<sub>2</sub>SO<sub>4</sub> 4.10<sup>-4</sup> M.
3. Si on mélange 250 ml d'une solution de Pb(NO<sub>3</sub>)<sub>2</sub> 1,6 10<sup>-3</sup> M avec 750 ml d'une solution de Na<sub>2</sub>SO<sub>4</sub> 2,4 10<sup>-3</sup> M.
4. On ajoute une solution de nitrate d'argent concentrée goutte à goutte à une solution aqueuse dont la concentration en ions Cl<sup>-</sup> est de 0,1M et la concentration en ions Br<sup>-</sup> est de 10<sup>-4</sup> M. Quelle est le sel qui précipité le premier AgCl ou AgBr ??

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**Scott C. Dulebohn**



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