

- 10 Magnesium metal reacts with oxygen to form magnesium oxide according to the following equation:



What mass of magnesium metal is required to produce 20.00 grams of magnesium oxide?

- (A) 6.03 grams
(B) 12.06 grams
(C) 24.12 grams
(D) 48.24 grams

- 11 Use the table of solubility rules to answer Questions 11 and 12.

<i>Solubility of ionic compounds</i>	
<i>Soluble</i>	<i>Insoluble</i>
All Group 1 and NH_4^+ compounds	Phosphates (except Group 1 and NH_4^+)
All nitrates	Hydroxides (except Group 1 and NH_4^+ , Ba^{2+} , Ca^{2+} , Sr^{2+})
Chlorides (except Ag^+ , Pb^{2+})	
Sulfates (except Ag^+ , Pb^{2+} , Ba^{2+} , Ca^{2+})	

Which pair of solutions would produce a precipitate?

- (A) Lithium phosphate and ammonium nitrate
(B) Barium nitrate and silver chloride
(C) Magnesium sulfate and calcium hydroxide
(D) Ammonium chloride and strontium hydroxide

- 12 Which ionic equation correctly shows the formation of a precipitate?

- (A) $\text{BaCl}_{2(aq)} + \text{MgSO}_{4(aq)} \rightarrow \text{BaSO}_{4(s)} + \text{MgCl}_{2(aq)}$
(B) $\text{Zn}^{2+}_{(aq)} + \text{SO}_4^{2-} \rightleftharpoons \text{ZnSO}_{4(s)}$
(C) $\text{Mg}^{2+}_{(aq)} + \text{PO}_4^{3-} \rightleftharpoons 2\text{Mg}(\text{PO}_4)_{(s)}$
(D) $3\text{Zn}^{2+}_{(aq)} + 2\text{PO}_4^{3-}_{(aq)} \rightleftharpoons \text{Zn}_3(\text{PO}_4)_{2(s)}$

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