

**Intro to Stoichiometry – Moles to Grams NOTES**

**Ex. 1:** Given 2.5 moles of  $\text{N}_2$ , how many grams of  $\text{NH}_3$  would be produced?

$$\frac{2.5 \text{ moles } \text{N}_2}{1} \times \frac{2 \text{ moles } \text{NH}_3}{1 \text{ mole } \text{N}_2} \times \frac{17.04 \text{ grams } \text{NH}_3}{1 \text{ mole } \text{NH}_3}$$

Answer: 85.2 grams of  $\text{NH}_3$

**Ex. 2:** Given 4.5 moles of  $\text{H}_2$ , how many grams of  $\text{N}_2$  will be used?

$$\frac{4.5 \text{ moles } \text{H}_2}{1} \times \frac{1 \text{ moles } \text{N}_2}{3 \text{ mole } \text{H}_2} \times \frac{28.01 \text{ grams } \text{N}_2}{1 \text{ mole } \text{N}_2}$$

Answer: 42.02 grams of  $\text{N}_2$

**Ex. 3:** Given 12.3 grams of  $\text{NH}_3$ , how many moles of  $\text{N}_2$  were needed?

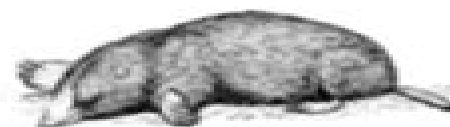
$$\frac{12.3 \text{ grams } \text{NH}_3}{1} \times \frac{1 \text{ mole } \text{NH}_3}{17.04 \text{ grams } \text{NH}_3} \times \frac{1 \text{ mole } \text{N}_2}{2 \text{ moles } \text{NH}_3}$$

Answer: 0.36 moles of  $\text{N}_2$

**Ex. 4:** Given 56 grams of  $\text{N}_2$ , how many moles of  $\text{NH}_3$  were produced?

$$\frac{56 \text{ grams } \text{N}_2}{1} \times \frac{1 \text{ mole } \text{N}_2}{28.01 \text{ grams } \text{N}_2} \times \frac{2 \text{ moles } \text{NH}_3}{1 \text{ mole } \text{N}_2}$$

Answer: 4.00 moles of  $\text{NH}_3$



# The Mole Hole Game Worksheet Chemistry

**R Pring**



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