Marrie	Date	Period
RATE CONCEPT: WAS	einforcement Worksheet – Propris unique preparties allow life to exist a	erties of Water
The structure of the wat means that it has a regio positive charge (the hyd form hydrogen bonds. A a slightly negative atom • High specific I heat energy to in • Cohesion: The to hydrogen bon • Adhesion: The	or inclicate gives water unique properti- a with a stight negative charge (the oxy- reges atoms). The oppositely charged in a hydrogen band is an attraction between Hydrogen bands are responsible for se- beat. Water resists charges in neoperate crease in temperature.  a attraction among molecules of a subst- ds makes water molecules "stick" togetle attraction among molecules of different to many other materials because of hy	es. Water is a gotar molecule, which gen acons), and a region with a slight agions of water trolecules immuct to a slightly positive hydrogen atom and veral important properties of water, sec it crust absorb a large amount of sace is called cohesion. Cohesion due her.
interiors, and chemical s another, a solution is for Substances that are pres	ne important for life dissolve in water. ' eactions in the cell take place in this wa reed. The substance present is the great ent in lower amounts and dissolve in the dissolve polar molecules and ions.	der. When one substance dissolves in lest amount is called the solvent.
hydrogen ion (a proton) accept, hydrogen ions. / scale. An acid has a low pill (pH above 7) and a l	dissolve in water they break up into ione when it dissolves in water is an acid. If I solution's acidity, or its hydrogen ion pH (pH below 7) and a high hydrogen ow hydrogen ion concentration. Organic discupt many biological processes.	lases are correpounds that remove, or concentration, is measured on the pH ion concentration. A base has a high
Explain why water is	a potar molocule?	
2. How do bydrogen ba	ada form?	
3. What are the two part	s of a solution?	
4. What types of molecu	ties dissolve easily in water? What type	s do not dissolve easily in water?
5. What does pH mean?		

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