

PRE-LABORATORY QUESTIONS
TITRATION OF A DIPROTIC ACID

Name _____

Laboratory Instructor _____ Day and time lab meets _____

1. What is a diprotic acid? (1 pt)
2. Give the balanced chemical reaction for the titration of oxalic acid with NaOH. (1 pts)



3. When titrating 25.0 mL of 0.10 M H_2SO_4 with 0.10 M NaOH, how many mL of NaOH will you have added to reach the first equivalence point? How many mL of NaOH will you have added to reach the second equivalence point? Show calculations. (4 pts)
4. A student completes a titration of an unknown diprotic acid. In their experiment, 0.79 g of the acid was dissolved in 250.0 mL of DI water. It required 13.48 mL of 1.0 M NaOH to reach the second equivalence point. What is the molar mass of the acid? Looking at the list of possible unknowns, what is likely to be the identity of the compound? Show calculations. (4pts)

Redox Titration Pre Lab Questions Answers

**Antonín Berka, Jaroslav
Vulterin, Jaroslav Zýka**



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Table of Contents Redox Titration Pre Lab Questions Answers

1. Understanding the eBook Redox Titration Pre Lab Questions Answers
 - The Rise of Digital Reading Redox Titration Pre Lab Questions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Redox Titration Pre Lab Questions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Redox Titration Pre Lab Questions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Redox Titration Pre Lab Questions Answers
 - Personalized Recommendations
 - Redox Titration Pre Lab Questions Answers User Reviews and Ratings
 - Redox Titration Pre Lab Questions Answers and Bestseller Lists
5. Accessing Redox Titration Pre Lab Questions Answers Free and Paid eBooks
 - Redox Titration Pre Lab Questions Answers Public Domain eBooks
 - Redox Titration Pre Lab Questions Answers eBook Subscription Services
 - Redox Titration Pre Lab Questions Answers Budget-Friendly Options
6. Navigating Redox Titration Pre Lab Questions Answers eBook Formats

- ePub, PDF, MOBI, and More
- Redox Titration Pre Lab Questions Answers Compatibility with Devices
- Redox Titration Pre Lab Questions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Redox Titration Pre Lab Questions Answers
 - Highlighting and Note-Taking Redox Titration Pre Lab Questions Answers
 - Interactive Elements Redox Titration Pre Lab Questions Answers
- 8. Staying Engaged with Redox Titration Pre Lab Questions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Redox Titration Pre Lab Questions Answers
- 9. Balancing eBooks and Physical Books Redox Titration Pre Lab Questions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Redox Titration Pre Lab Questions Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Redox Titration Pre Lab Questions Answers
 - Setting Reading Goals Redox Titration Pre Lab Questions Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Redox Titration Pre Lab Questions Answers
 - Fact-Checking eBook Content of Redox Titration Pre Lab Questions Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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