

Relative Molecular Mass

The Relative Molecular Mass (M_r) is the sum of the Relative Atomic Mass (A_r) of each atom in the molecule.

How to find the relative molecular mass?

1. Determine the molecular formula of the molecule.
2. Determine the number of atoms of each element in the molecule.
3. Use the periodic table to determine the atomic mass of each element.
4. Multiply the number of atoms of each element with the atomic mass.
5. Add the values to get the relative molecular mass.

Example:

Find the M_r of ammonia

Molecular formula of ammonia is NH_3

$$\text{N} = 1 \times 14 = 14$$

$$\text{H} = 3 \times 1 = 3$$

$$M_r \text{ of ammonia} = 14 + 3 = 17$$

A_r of N = 14

A_r of H = 1

Example:

Find the M_r of ammonium sulfate

Molecular formula is $(\text{NH}_4)_2\text{SO}_4$

$$\text{N} = 2 \times 14 = 28$$

$$\text{H} = 4 \times 2 \times 1 = 8$$

$$\text{S} = 1 \times 32 = 32$$

$$\text{O} = 4 \times 16 = 64$$

$$M_r \text{ of } (\text{NH}_4)_2\text{SO}_4 = 28 + 8 + 32 + 64 = 132$$

A_r of N = 14

A_r of H = 1

A_r of S = 32

A_r of O = 16

Relative Atomic Mass And Chemical Formulas

Peter B. Moore



Relative Atomic Mass And Chemical Formulas:

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historical perspective showing in outline how important themes of chemistry have developed
Weights, Molecular Formulas, Microcosmic Magnitudes Leonard Kollender Nash, 1966

Stoichiometry: Atomic
An Introduction to

Chemistry Peter B. Moore, 2020-02-24 This publication is an entry level textbook designed to meet the needs of college students who learned some chemistry in their high school years but not enough to prepare them for advanced courses in chemistry or to satisfy the chemistry prerequisite for courses they might want to take in other scientific disciplines The history of chemistry is emphasized to an unusual degree here primarily to give the narrative a storyline but its historical emphasis has an important secondary benefit Much of the vocabulary chemists use to describe chemical phenomena today emerged early in the development of the discipline when their understanding of them was still in a primitive state As such the persistence of these words and the concepts behind them makes sense only in the light of history
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introduces the basic theories and concepts of a particular topic focusing on relevant equations Worked examples illuminate each type of problem with carefully explained step by step solutions Since chemistry problem can be presented in a number of ways the examples include several versions of each questions To help students understand and retain the procedures the solutions discuss not only what steps to carry out to reach solutions but why The second edition contains additional problems at the end of each chapter with varying degrees of difficulty and many of the original examples have been revised Book jacket

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General and Inorganic Chemistry in Medicine Budin Michov, 2024-12-09 This book is intended for students in medicine pharmacy and dentistry physicians dentists pharmacists biochemists and more In General Chemistry the laws of chemistry the structure of simple and complex compounds chemical bonds solutions chemical reactions kinetics equilibrium thermodynamics protolytic and redox processes and sorption are discussed In Inorganic Chemistry chemical elements inorganic compounds and their significance for medicine are presented It is focused on developing metal based diagnostic and therapeutic agents The significance of coordination chemistry to modulate enzyme activity is discussed The production of reactive oxygen species selectively damaging cancer cells is described too Short biographies of chemists and scientists which have rendered services to general and inorganic chemistry in medicine are given in a person index

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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