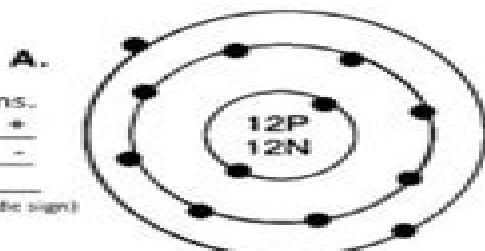
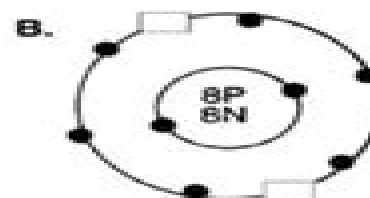


# Types of Bonds

Look at the Bohr Models below, move the electrons from one model to the other as they would if these two atoms formed a bond.



After moving the electrons.  
How many Protons \_\_\_\_ +  
How many Electrons \_\_\_\_ -  
What is the difference \_\_\_\_  
(include sign)



After moving the electrons.  
How many Protons \_\_\_\_ +  
How many Electrons \_\_\_\_ -  
What is the difference \_\_\_\_

1. If atom A loses electrons to atom B,
  - a. how many electrons will atom A lose? \_\_\_\_\_
  - b. how many electrons will atom B gain? \_\_\_\_\_
  - c. what will be the oxidation number of atom A? \_\_\_\_\_ (include sign)
  - d. what will be the oxidation number of atom B? \_\_\_\_\_ (include sign)
  - e. what will be the total charge of the compound formed? Number format.  
\_\_\_\_\_

f. what type of bond will form?  
\_\_\_\_\_

2. Explain why an element's oxidation number is related to the group on the periodic table to which it belongs.

For Families 1, 2, 3, and 4 the oxidation number is \_\_\_\_\_

For Families 5, 6, 7, and 8 the oxidation number is \_\_\_\_\_

**Directions:** Complete the table comparing ionic compounds and covalent compounds.

| Characteristic                     | Ionic compounds                                   | Covalent compounds                                |
|------------------------------------|---------------------------------------------------|---------------------------------------------------|
| 3. How the compound is formed      | electrons are shared<br>electrons are transferred | electrons are shared<br>electrons are transferred |
| 4. Smallest particle               | proton, neutron, electron                         | proton, neutron, electron                         |
| 5. Usual state at room temperature | Solid, Liquid, Gas                                | Solid, Liquid, Gas                                |

# What Are Ionic Bonds Worksheet 233 Physical Science

**Karin Nielsen-Saines**



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