

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Regioselectivity and Stereochemistry

Bond Types, Strength, Length, Polarity and Dipole Moment

Types of Bonding Interactions

There are two types of bonding: ionic bonding and covalent bonding.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron).



- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Strength of Bonds



Bond and Reaction Rates



Bond Energy



There are two types of bonds:

1. **Polar bonds**



2. **Nonpolar bonds**



Sigma and pi bonds

Covalent bonds are classified as sigma and pi bonds.

- **Sigma bonds** are formed by head-on overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **Single bonds** are sigma bonds.
- **Double bonds** are sigma and pi bonds.



pi bonds are weaker than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in electronegativity, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar).
- If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar).
- If the difference in electronegativity is greater than 1.7, the electrons are not shared (ionic).

- A difference of 1.7 or higher indicates that the electrons are not shared (ionic).

- A difference of 1.7 or higher indicates that the electrons are not shared (ionic).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.

- The more electronegative atom has a partial negative charge (δ⁻) and the less electronegative atom has a partial positive charge (δ⁺).

- **δ⁻** (delta minus) indicates an electron-deficient atom.

- **δ⁺** (delta plus) indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if the dipoles are present. A good example is carbon dioxide.



Carbon dioxide is a linear molecule and the dipoles cancel.

Bond Length and Bond Strength

The longer the bond, the weaker it is.



Double bonds are stronger than single bonds.

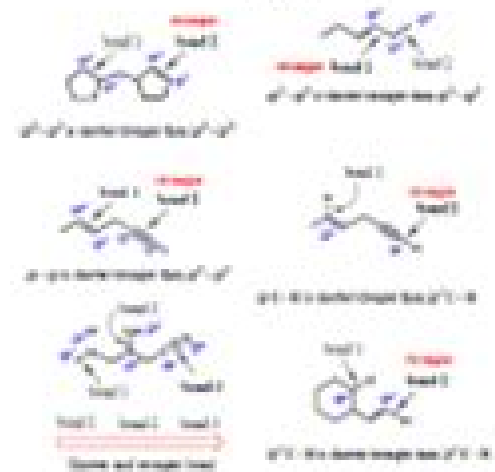


The longer the bond, the weaker it is.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp ³ -C	sp ² -C	sp-C
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	47 kcal/mol	58 kcal/mol
Hybridization	sp ³ -sp ³	sp ² -sp ²	sp-sp
C-H Bond	sp ³ -H	sp ² -H	sp-H
Length	1.09 Å	1.08 Å	1.06 Å
Strength	98 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character means stronger bond

Examples



Organic Chemistry Guide

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