

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Two types of covalent bonds

There are two types of covalent bond: ionic and covalent bond.

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron affinity).



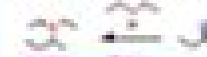
- **Covalent bonds** are formed by sharing of electrons between two atoms.



The difference between ionic and covalent bonds



Bond and Electron sharing



Bonding energy



There are two types of covalent bonds



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

- **Sigma bonds** are formed by head-on overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **Single bond** is a sigma bond.
- **Double bond** consists of one sigma and one pi bond.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be polar and nonpolar.



- If the difference in electronegativity is less than 0.5, the electron is shared equally (nonpolar bond).
- If the difference in electronegativity is between 0.5 and 1.7, the electron is shared unequally (polar bond).
- If the difference in electronegativity is greater than 1.7, the electron is transferred from one atom to another (ionic bond).

A difference of 1.7 or higher can be high for the electron to be shared and a lower bond is formed.

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).



- The dipole moment is defined as a vector pointing from the less electronegative to the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).
- δ^- (delta minus) indicates an electron deficient atom.
- δ^+ (delta plus) indicates an electron rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moment.

The polarity of the molecule is determined by the vector sum of the dipole moments of the bonds. A good example is carbon dioxide.



Carbon dioxide is a nonpolar molecule because the vector sum of the dipole moments is zero.

Bond Length and Bond Strength

The length of the bond is measured in Å.

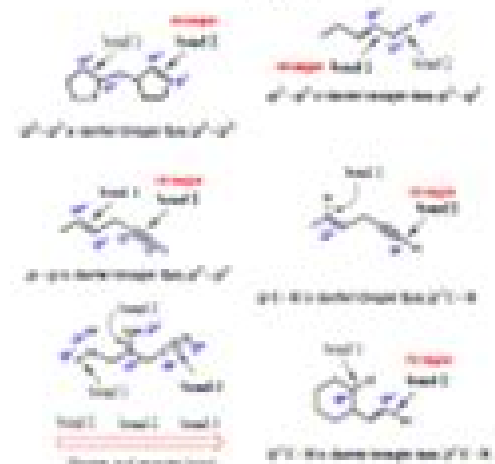


The length of the bond is measured in Å.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.54 Å	1.34 Å	1.20 Å
Strength	33 kcal/mol	112 kcal/mol	200 kcal/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	1.09 Å	1.08 Å	1.06 Å
Strength	98 kcal/mol	110 kcal/mol	120 kcal/mol

Higher s character means stronger bond

Examples



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Kathleen Armour



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