

# Working a Stoichiometry Problem

6.50 grams of aluminum reacts with an excess of oxygen. How many grams of aluminum oxide are formed?



$$6.50 \times 2 \times 101.96 \div 26.98 \div 4 = 12.3 \text{ g Al}_2\text{O}_3$$

<del>6.50 g Al</del>	<del>1 mol Al</del>	<del>2 mol Al<sub>2</sub>O<sub>3</sub></del>	<del>101.96 g Al<sub>2</sub>O<sub>3</sub></del>
	<del>26.98 g Al</del>	<del>4 mol Al</del>	<del>1 mol Al<sub>2</sub>O<sub>3</sub></del>

$$= ? \text{ g Al}_2\text{O}_3$$

# Sample Problem Set Stoichiometry Chemistry

**Arihant Experts**



### **Sample Problem Set Stoichiometry Chemistry:**

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