

Physical Science EOCT Study Guide – Domain 3

Physical Energy, Force, and Motion

Standards: PS.3-7 and 8

1. Students will understand relationships and flow of energy within a system.
 - a. Explain thermal energy (heat) and kinetic energy (movement) and how they are related.
 - b. Explain heat. Thermal energy flows from objects that are hotter to objects that are cooler.
 - c. Explain specific heat. The amount of energy needed to raise the temperature of 1 kg of a substance by 1°C.
 - d. What is the formula for determining that change in heat that energy can add?

$$Q = mc\Delta T \quad \Delta T = T_2 - T_1$$
 - e. A sample of 0.5 kg of water at 20°C and specific heat of 4186 J/kg°C. Find the change in heat energy if the block is 2 mmol from 10°C to 20°C.

$$Q = mc\Delta T \quad 20 = 20 - 10 = 10 \quad 5000 \text{ J and } 5000 \text{ J}$$
 - f. Explain conduction, convection, and radiation. Give examples of each.

Conduction is a transfer of energy by contact. Example: a metal spoon in a pot of soup.

Convection is a transfer of energy by movement of a fluid. Example: a pot of soup.

Radiation is a transfer of energy by electromagnetic waves. Example: a microwave oven.

2. Explain the following phase change diagrams.



3. Students will determine relationships among force, mass and motion.
 - a. Explain the difference between distance and displacement.

Distance is the total path length. Displacement is the change in position.
 - b. What is the difference between average speed and instantaneous speed?

$$\text{Average Speed} = \frac{\text{Total Distance}}{\text{Total Time}}$$

$$\text{Instantaneous Speed} = \frac{d}{dt}$$
 - c. What is the difference between acceleration and force?

$$F = ma$$

$$v = \frac{d}{t}$$

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it She presents an array of tested activities and assignments through which students can progressively reflect on monitor and improve their learning skills describes how they can be integrated with different course components and on various schedules and elucidates how to intentionally and seamlessly incorporate them into course design to effectively meet disciplinary and student development objectives Recognizing that most faculty are unfamiliar with these strategies she also recommends how to prepare for introducing them into the classroom and adding more as instructors become more confident using them The book concludes with descriptions of courses from different fields to offer models and ideas for implementation At a time of so much concern about what our students are learning in college and how well prepared they are for the challenges of tomorrow s economy and society self regulated learning provides a reassuring solution particularly as studies indicate that struggling students benefit the most from practicing it *Essentials of Physical Chemistry* Don Shillady,2011-07-27 At a time when U S high school students are producing low scores in mathematics and science on international examinations a thorough grounding in physical chemistry should not be considered optional for science undergraduates Based on the author s thirty years of teaching *Essentials of Physical Chemistry* merges coverage of calculus with chemistry and molecular physics in a friendly yet thorough manner Reflecting the latest ACS guidelines the book can be used as a one or two semester course and includes special topics suitable for senior projects The book begins with a math and physics review to ensure all students start on the same level and then discusses the basics of thermodynamics and kinetics with mathematics tuned to a level that stretches students abilities It then provides material for an optional second semester course that shows students how to apply their enhanced mathematical skills in a brief historical development of the quantum mechanics of molecules Emphasizing spectroscopy the text is built on a foundation of quantum chemistry and more mathematical detail and examples It contains sample classroom tested exams to gauge how well students know how to use relevant formulas and to display successful understanding of key concepts Coupling the development of mathematical skills with chemistry concepts encourages students to learn mathematical derivations Mini biographies of famous scientists make the presentation more interesting from a people point of view Stating the basic concepts of quantum chemistry in terms of analogies provides a pedagogically useful technique Covering key topics such as the critical point of a van der Waals gas the Michaelis Menten equation and the entropy of mixing this classroom tested text highlights applications across the range of chemistry forensic science pre medical science and chemical engineering In a presentation of fundamental topics held together by clearly established mathematical models the book supplies a quantitative discussion of the merged science of physical chemistry **ASE Automotive Service Excellence A1-A8 exam Study Guide with 500 Questions** Alisa Utomo,Heidi Schneider,Trent Vanotti,2016-02-01 Comprehensive ASE A1 A8 exam questions Covers the following A1 Auto Engine Repair A2 Auto Automatic Transmission Transaxle A3 Auto Manual Drive Train A4 Auto Suspension A5 Auto Brakes A6 Auto Electrical Electronic Systems A7 Auto Heating A8 Auto Engine Performance You have the Edge You now have an

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