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## Section 2-2 Properties of Water (pages 40-43)

## Key Concepts

- Why are water molecules polar?
- · What are acidic solutions? What are basic solutions?

#### The Water Molecule (pages 40-41)

- Is the following sentence true or false? A water molecule is neutral.
- Why is a water molecule polar? There is an uneven distribution of electrons between the oxygen and hydrogen atoms.
- Circle the letter of each sentence that is true about hydrogen bonds.
  - a. A hydrogen bond is stronger than an ionic bond.
  - (b) The attraction between the hydrogen atom on one water molecule and the oxygen atom on another water molecule is an example.
  - c. A hydrogen bond is stronger than a covalent bond.
  - (d.) They are the strongest bonds that form between molecules.
- 4. Complete the table about forms of attraction.

#### FORMS OF ATTRACTION

Form of Attraction	Definition
Cohesion	Attraction between molecules of the same substance
Adhesion	Attraction between molecules of different substances

#### Solutions and Suspensions (pages 41-42)

- What is a mixture? A mixture is a material composed of two or more elements or compounds that are physically mixed together but not chemically combined.
- 7. The greatest solvent in the world is \_\_\_\_\_\_\_
- What is a suspension? A suspension is a mixture of water and nondissolved substances that are so small they do not settle out.



# **Section 2 2 Properties Of Water Worksheet Answers**

Jay K. Hackett

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