

12

STOICHIOMETRY

SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353–358)

This section explains how to calculate the masses of reactants required or products formed in a stoichiometric process. It teaches you how to interpret chemical equations in terms of measuring moles, representative particles, masses, and gas volume at STP.

► Using Everyday Equations (pages 353–354)

1. How can you determine the quantities of reactants and products in a chemical reaction?

You must enter **2** or **balancing** to qualify.

2. Quantity usually means the _____ **amount** _____ of a substance expressed in grams or moles.

3. A hot-dogger is no the fault ofers 3 wheels (3H), 2 sides boards (2B), 1 top (T), 1 base (B), and 4 legs (L). Write a “balanced equation” for the construction of this hot-dogger.

$3\text{H} + 2\text{B} + 1\text{T} + 1\text{B} + 4\text{L} = 1\text{Hot-Dogger}$

► Using Balanced Chemical Equations (page 354)

4. In the following sentence (or on your Student Laboratory) is the calculation of quantities in chemical reactions. _____ **True**

5. Calculations using balanced equations are called _____ **stoichiometric calculations**

► Interpreting Chemical Equations (pages 354–357)

6. Does what element is ammonia produced? How is it used?

Ammonia NH_3 is a compound of nitrogen and hydrogen. It is used as a fertilizer.

7. Circle the letter of the term that tells what kind of information you **CANNOT** get from a chemical equation.

- | | |
|-----------------------------|------------------------|
| a. moles | d. volume |
| b. mass | e. number of particles |
| c. size of particles | |

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L Darling-Hammond



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