

Section 12.3 Limiting Reagent & Percent Yield

Answer Key

Michelle Galian

Vocabulary

1. e
2. c
3. d
4. a
5. b

1. Answer: 127g Cu
Because...

Find the limiting reagent:

$$54\text{g Al} \times 1 \text{ mole Al}/27\text{g Al} = 2 \text{ mole Al}$$

$$319\text{g CuSO}_4 \times 1 \text{ mole CuSO}_4/159.5\text{g CuSO}_4 = 2 \text{ mole CuSO}_4 \text{ (limiting reagent)}$$

$$2 \text{ mole CuSO}_4 \times 3 \text{ mole Cu}/3 \text{ mole CuSO}_4 \times 63.5\text{g Cu}/1 \text{ mole Cu} = 127\text{g Cu}$$

(Multiply the founded limiting reagent with moles of Copper (3) over moles of CuSO₄ (3) from the equation to cancel out moles of CuSO₄.

Then, to cancel out 3 moles Cu, multiply it with grams Cu (because that is also what the question asks for) over 1 mole Cu. After canceling and multiplying out what is left, you come to 127g Cu.)

2. $27\text{g Cu} \times 1 \text{ mole Cu}/63.5\text{g Cu} \times 2 \text{ mole Cr}/3 \text{ mole Cu} \times 52\text{g Cr}/1 \text{ mole Cr} = 14.74\text{g Cr}$

("excess of CuSO₄" is irrelevant to the problem you are solving, so you can cross that out. You are left only to work with 27g Cu, and need to find grams Cr. So you divide 27g Cu by its molar mass [from periodic table] and then multiply it by the mole to mole ratio of 2 moles Cr over 3 moles Cu. You put 3 moles Cu on the bottom to cancel it out with 1 mole Cu. Lastly, to cancel out 2 moles Cr, multiply it with the molar mass [52g] over 1 mole Cr and multiply out what is uncrossed.)

Section 12 3 Limiting Reagent And Percent Yoeld

Morris Hein,Susan Arena,Cary Willard

Section 12 3 Limiting Reagent And Percent Yoeld:

Chemistry: The Central Science Theodore L. Brown,H. Eugene LeMay Jr.,Bruce E. Bursten,Catherine Murphy,Patrick Woodward,Steven Langford,Dalius Sagatys,Adrian George,2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation OAT 2017-2018 Strategies, Practice & Review with 2 Practice Tests Kaplan Test Prep,2016-10-04 Issued with 16 pages of detachable study sheets and access to two full length practice tests

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