Chemistry Unit 15: Chemical Equilibrium Worksheet

Practice 15.1 Writing Equilibrium Equations

Write the equilibrium expressions (K_{oo}) for the following reactions

- 2CO(g) ← C(s) + CO₂(g)
- N₂O_b(g) ↔ 2NO₂(g).
- 2lCl(g) ↔ I₂(g) + Cl₂(g),
- Fe₂O₃(s) + 3H₂(g) ↔ 2Fe(s) + 3H₂O(g)
- 2CO(g) + 4H₂O(g) ↔ 3O₂(g) + 2CH₄(g)
- 6. $P_d(s) + 6NO(g) \leftrightarrow P_dO_d(s) + 3N_2(g)$

15.2 Calculating Equilibrium Constants

Be sure to write out the equilibrium (K.,) expression for each problem before trying to solve it!

1. What is the equilibrium expression for the following reaction?

$$SmO_2(s) + 2CO(g) \leftrightarrow Sm(s) + 2CO_2(g)$$

- At equilibrium the following concentrations are recorded: 20.0 M of H₂, 18.0 M of CO₂, 12.0 M H₂O, and 5.9 M of CO at 427 °C. What is the value of K₆₀₀ at this temperature for the following reaction? CO₂(g) + H₂(g) → CO(g) + H₂O(g)
- At equilibrium a mixture of N₂, H₂, and NH₃ gas at 500°C is determined to consist of 0.602 M N₂, 0.420 M H₂, and 0.113 M of NH₃. What is the equilibrium constant at this temperature? Does this equilibrium favor the reactants or the products? N₂(g) + 3H₂(g) ← 2NH₃(g)
- Given the equilibrium concentrations of 0.15 M NO, 0.12 M H₂ and 1.6 M N₂, Calculate the equilibrium constant for the following reaction. Does this reaction favor the reactants or the products?
 2NO(g) = 2H₂(g) ↔ N₂(g) = 2H₂O(l)
- What is the equilibrium constant for the following reaction if the equilibrium concentrations for the gases are 0.3 M for C₂H₆, 0.6 M for O₂, 1.3 M for CO₂ and 1.8 M for H₂0? Does the reaction favor the reactions or products?
 2C₂H₆(g) + 7O₂(g) 4CO₂(g) + 6H₂O(g)
- Nitrogen dioxide reacts to form dinitrogen tetraoxide. What is the equilibrium concentration of N₂O_n if the K_{eq} = 4.70 and the equilibrium concentration of NO₂ is 0.0310 M?
 2NO₂(g) → N₂O₄(g)

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