## 2.1.4 & 2.1.5 - Acids and redox

What are standard solutions?	A solution for which its concentration is known accurately.
What 3 things should you ensure when making standard solutions?	Ensure you have clean dry equipment.     Make washings to remove any salt <u>BEFORE MAKING</u> up the solution to the specified amount.     Invert the volumetric flask to ensure everything is dissolved.
How does ammonia react with acid?	<ul> <li>ammonia + acid → ammonium salt</li> <li>NH<sub>3</sub>(aq) + HNO<sub>3</sub> (aq) → NH<sub>4</sub>NO<sub>3</sub>(aq)</li> </ul>
What are strong and weak acids ? Give 3 strong acids and 1 weak acid	<ul> <li>Strong acids dissociate fully to release H' ions in a solution e.g HCl → H' + Cl'</li> <li>Weak acids dissociate partially to release H' ions in a solution – exist in equilibrium</li> <li>Strong acids include: HCl, H<sub>2</sub>SO<sub>4</sub>, and HNO<sub>3</sub>.</li> <li>Weak acids include carboxylic acids such as CH<sub>2</sub>COOH.</li> </ul>
What are strong and weak bases? Give 1 weak base	<ul> <li>Strong bases dissociate fully to release OH ions in a solution</li> <li>Weak bases dissociate partially to release OH ions in a solution – exist in equilibrium e.g. NH<sub>3</sub></li> </ul>
What are diprotic and triprotic acids (with examples) and what does this mean for neutralisation?	<ul> <li>Diprotic acids and triprotic acids donate two and three protons respectively.</li> <li>Examples include: H<sub>2</sub>SO<sub>4</sub> and H<sub>3</sub>PO<sub>4</sub>.</li> <li>They need 2x or 8x the number of moles of base to be neutralized.</li> </ul>

# Ocr Chemistry A2 Unofficial Mark Scheme 2014

Alisdair Thorpe, Dave Newton, Chris
Otter

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