

## Flame Test Kit Student Laboratory Kit

### Introduction

Just as a fingerprint is unique to each person, the color of light emitted by metals heated in a flame is unique to each metal. In this laboratory activity, the characteristic color of light emitted for calcium, copper, lithium, potassium, sodium, and strontium will be observed.

### Chemical Concepts

- Flame Tests
- Absorption/Emission

### Background

#### Absorption and Emission of Light in a Flame

When a substance is heated in a flame, the substance's electrons absorb energy from the flame. This absorbed energy allows the electrons to be promoted to excited energy levels. From these excited energy levels, the electrons naturally want to make a transition, or relax, back down to the ground state. When an electron makes a transition from a higher energy level to a lower energy level, a particle of light called a *photon* is emitted. A photon is commonly represented by a squiggly line (see Figure 1).

An electron may relax all the way back down to the ground state in a single step, emitting a photon in the process. Or, an electron may relax back down to the ground state in a series of smaller steps, emitting a photon with each step. In either case, the energy of each emitted photon is equal to the difference in energy between the excited state and the state to which the electron relaxes. The energy of the emitted photon determines the color of light observed in the flame. Because colors of light are commonly referred to in terms of their wavelength, equation 1 is used to convert the energy of the emitted photon to its wavelength.

$$\Delta E = \frac{hc}{\lambda} \quad \text{Equation 1}$$

In Equation 1,

$\Delta E$  is the difference in energy between the two energy levels in Joules,

$h$  is Planck's constant ( $h = 6.626 \times 10^{-34}$  J-sec),

$c$  is the speed of light ( $c = 2.998 \times 10^8$  m/sec), and

$\lambda$  is the wavelength of light in meters.

Wavelengths are commonly listed in units of nanometers ( $1 \text{ m} = 1 \times 10^9 \text{ nm}$ ), so a conversion between meters and nanometers is generally made.

The color of light observed when a substance is heated in a flame varies from substance to substance. Because each element has a different electronic configuration, the electronic transitions for a given substance are unique. Therefore, the difference in energy between energy levels, the exact energy of the emitted photon, and its corresponding wavelength and color are unique to each substance. As a result, the color observed when a substance is heated in a flame can be used as a means of identification.

#### The Visible Portion of the Electromagnetic Spectrum

Visible light is a form of electromagnetic radiation. Other familiar forms of electromagnetic radiation include  $\gamma$ -rays such as those from radioactive materials and from space, X-rays which are used to detect bones and teeth, ultraviolet (UV) radiation from

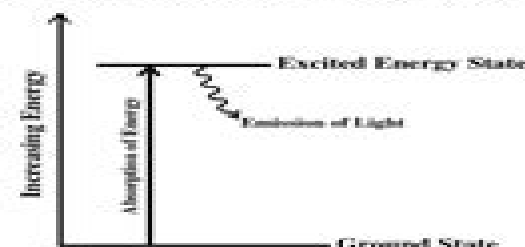


Figure 1. Absorption and Emission of Light.

# Sample High School Flame Test Lab With Answer Key

**S Ashworth**



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