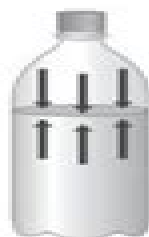


15. Describe the two diagrams of a bottled carbonated beverage below as *greater pressure* or *lower pressure*, and then as *greater solubility* or *lower solubility*. How do these two examples illustrate the relationship between the solubility of a gas and its vapor pressure?

In the closed bottle, the pressure is high and lots of  $\text{CO}_2$  is dissolved in the  
carbonated beverage. Once the cap is removed, the vapor pressure decreases,  
allowing more  $\text{CO}_2$  in the carbonated beverage to escape.



greater pressure and greater solubility      lower pressure and lower solubility

16. How does a solution become supersaturated? If the temperature of a solution having  
a small excess of solid solute is raised, the solute dissolves. If the solution is then  
allowed to cool slowly, the excess solute may stay dissolved at a temperature below  
the temperature at which it would ordinarily crystallize.

## SECTION 16.2 CONCENTRATIONS OF SOLUTIONS (pages 480–486)

*This section explains how to solve problems involving molarity of a solution, how to prepare dilute solutions from more concentrated solutions, and what is meant by percent by volume and percent by mass.*

### ► Molarity (pages 480–482)

1. A measure of the amount of solute dissolved in a given quantity of solvent is the concentration of a solution.
2. The most important unit of concentration in chemistry is molarity.
3. Is the following sentence true or false? Molarity is the number of moles of dissolved solute per liter of solvent. false

## Section 16 3 Colligative Properties Of Solutions Key

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General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition. Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works. This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value.

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