

12

STOICHIOMETRY

SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353–358)

This section explains how to calculate the masses of reactants required or products formed in a stoichiometric process. It teaches you how to interpret chemical equations in terms of measuring moles, representative particles, masses, and gas volume at STP.

► Using Everyday Equations (pages 353–354)

1. How can you determine the quantities of reactants and products in a chemical reaction?

You must enter **2** or **balancing** to qualify as

2. Quantity usually means the _____ **amount** _____ of a substance expressed in grams or moles.

3. A bookkeeper is on the fault lines 3 shelves (SF), 2 side boards (SB), 1 top (T), 1 base (B), and 4 legs (L). Write a “balanced equation” for the construction of this bookcase.

$3\text{SF} + 2\text{SB} + 1\text{T} + 1\text{B} + 4\text{L} = 1\text{Bookcase}$

► Using Balanced Chemical Equations (page 354)

4. In the following sentence (as in Robert Hooke’s discovery is the calculation of quantities in chemical reactions), _____ **is not** _____

5. Calculations using balanced equations are called _____ **stoichiometric calculations** _____

► Interpreting Chemical Equations (pages 354–357)

6. From what elements is ammonia produced? How is it used?

Ammonia combines with oxygen to produce nitrogen and hydrogen; it is used as a fertilizer.

7. Circle the letter of the term that tells what kind of information you **CANNOT** get from a chemical equation.

- | | |
|-----------------------------|------------------------|
| a. moles | d. volume |
| b. mass | e. number of particles |
| c. size of particles | |

Section 12 1 The Arithmetic Of Equations Practice Problems

Mark Zegarelli



Section 12 1 The Arithmetic Of Equations Practice Problems:

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